



# Sage 50erp User Guide

**Annelies Wilder-Smith**



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web during oxidation and reduction the oxidation numbers of atoms change if an atom is oxidized its oxidation number increases ie it becomes more ve or less ve if an atom is reduced its oxidation number decreases ie it becomes less ve or more ve

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web you should try to answer the questions without referring to your textbook if you get stuck try asking another group for help 1 identify the oxidation state of each element in the following a na 2 so 4 b mno 4 c  $\text{KClO}_3$  d  $\text{CH}_3\text{OH}$  2 identify the oxidation state of nitrogen in the following a no 3 b no 2 c no 2

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web write the oxidation and reduction reactions for each redox reaction the first one is done for you 23  $\text{Fe}^{2+} + \text{Co} \rightarrow \text{Co}^{2+} + \text{Fe}$   
oxidation  $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$  reduction  $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$  24 3  $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$  oxidation reduction 25  $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$  oxidation reduction 26  $\text{O}_2$

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web in each of the following redox reactions indicate the elements that undergo changes in oxidation number and the magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a  $\text{I}_2 + 5\text{O}_2 \rightarrow 2\text{IO}_5$   
 $5\text{CO} + \text{I}_2 \rightarrow 5\text{CO}_2 + \text{I}_2$  o  $5\text{C} + \text{O}_2 \rightarrow 5\text{CO}_2$

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web jun 28 2010 answers i determine what is oxidized and what in each reaction identify the oxidizing agent and the reducing agent also 2  $\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$   $\text{Sr} \rightarrow \text{Sr}^{2+}$  oxidized reducing agent  $\text{O}_2 \rightarrow \text{O}^{2-}$  reduced ox ag 2  $\text{Li} + \text{S} \rightarrow \text{Li}_2\text{S}$   $\text{Li} \rightarrow \text{Li}^+$  oxidized red ag  $\text{S} \rightarrow \text{S}^{2-}$  reduced ox ag 3  $2\text{Cs} + \text{Br}_2$

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