

16. The dissociation of water is shown below:



(continued)

At 25 °C, the auto product of water,  $K_w$ , is  $1.0 \times 10^{-14}$ .

At 25 °C, the auto product of water,  $K_w$ , is  $1.0 \times 10^{-14}$ .

At 25 °C, the auto product of water,  $K_w$ , is  $1.0 \times 10^{-14}$ .

17. Explain what the dissociation of water is an indication of water's behavior.

Endothermic

18. Predict, using a calculation, whether a solution of 1.0 M  $HCl$  is acidic, basic or neutral.

Acidic

Explain by calculating pH.

19.  $pK_a$ ,  $pH$  and  $pOH$  are significant values.

Calculate  $pK_a$  of  $HCl$ .

Give your answer to two decimal places.

$$-\log(4.71 \times 10^{-4}) = 3.03$$

$$pK_a = 3.03$$

20. A solution of hydrochloric acid is shown with water and the resulting solution is shown at 25 °C.

Calculate the pH of the resulting solution at 25 °C.

Give your answer to two decimal places.

$$pH = 1.30$$

pH = \_\_\_\_\_ (1)

(Total: 10)

# Ocr Chemistry A2 Unofficial Mark Scheme 2014

**Scott C. Dulebohn**



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