

Part 2

GCSE Chemistry OCR A (Combined Science - **HIGHER**) Paper 2 Exam Walkthrough

Combined Science Higher - Chemistry

Walkthrough Mark Paper

12
(5)

Look at the equation for the equilibrium reaction between sulfur dioxide, SO_2 , oxygen and sulfur trioxide, SO_3 , in a closed system.



- (a) Predict the effect of adding more oxygen to the equilibrium mixture.

Explain your answer.

(2)

- (b) Predict the effect of increasing the total pressure in the equilibrium mixture.

Explain your answer.

(2)

- (c) Predict the effect of increasing the **temperature** of the equilibrium mixture.

Explain your answer.

(2)

- (d) Sulfur dioxide for this reaction is made by burning sulfur.



Calculate the mass of sulfur needed to make 48 tonnes of sulfur dioxide.

Give your answer to 2 significant figures.

The relative atomic mass, A_r , of S is 32.1 and of O is 16.0.

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Terry C. Jones



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Table of Contents Ocr Chemistry Papers Gcse

1. Understanding the eBook Ocr Chemistry Papers Gcse
 - The Rise of Digital Reading Ocr Chemistry Papers Gcse
 - Advantages of eBooks Over Traditional Books
2. Identifying Ocr Chemistry Papers Gcse
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Ocr Chemistry Papers Gcse
 - User-Friendly Interface
4. Exploring eBook Recommendations from Ocr Chemistry Papers Gcse
 - Personalized Recommendations
 - Ocr Chemistry Papers Gcse User Reviews and Ratings
 - Ocr Chemistry Papers Gcse and Bestseller Lists
5. Accessing Ocr Chemistry Papers Gcse Free and Paid eBooks
 - Ocr Chemistry Papers Gcse Public Domain eBooks
 - Ocr Chemistry Papers Gcse eBook Subscription Services
 - Ocr Chemistry Papers Gcse Budget-Friendly Options

6. Navigating Ocr Chemistry Papers Gcse eBook Formats
 - ePub, PDF, MOBI, and More
 - Ocr Chemistry Papers Gcse Compatibility with Devices
 - Ocr Chemistry Papers Gcse Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Ocr Chemistry Papers Gcse
 - Highlighting and Note-Taking Ocr Chemistry Papers Gcse
 - Interactive Elements Ocr Chemistry Papers Gcse
8. Staying Engaged with Ocr Chemistry Papers Gcse
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Ocr Chemistry Papers Gcse
9. Balancing eBooks and Physical Books Ocr Chemistry Papers Gcse
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Ocr Chemistry Papers Gcse
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Ocr Chemistry Papers Gcse
 - Setting Reading Goals Ocr Chemistry Papers Gcse
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Ocr Chemistry Papers Gcse
 - Fact-Checking eBook Content of Ocr Chemistry Papers Gcse
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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