

CBSE Class 9 Maths Sample Paper Set 6

UWDLGEV<OCVJGOCVKEU"
ENCUU<KZ"

OCZFOCTMU"<:2"
FWTCVIQP"<5"JTU"

General Instruction:

- All questions are compulsory.
- This question paper contains 30 questions divided into four Sections A, B, C and D.
- Section A comprises of 6 questions of 1 mark each. Section B comprises of 6 questions of 2 marks each. Section C comprises of 10 questions of 3 marks each and Section D comprises of 8 questions of 4 marks each.
- There is no overall choice. However, an internal choice has been provided in two questions in 1 mark each, two questions in 2 marks each, four questions of 3 marks each and three questions of 4 marks each. You have to attempt only one of the alternatives in all such questions.
- Use of Calculators is not permitted.

SECTION – A

Questions 1 to 6 carry 1 mark each.

- What will be the ratio of areas of the two parallelogram which lie on same base and between the same parallels?
- Give one example each of a binomial of degree 35 and of a monomial of degree 100.
- In which quadrant points (1, -1), (2, -2), (4, -5) and (3, -4) lie?
- Simplify: $\sqrt{72} + \sqrt{800} - \sqrt{18}$

OR

Simplify: $343^{\frac{1}{3}}$

- Find the value of x, so that AOB is a line.



- The following data have been arranged in ascending order of their values. 19, 20, 36, 46, 50, $x - 2$, $x + 4$, 59, 70, 80, 83, 88. If the median is 57 find x.

OR

Find the mean of the factors of 24.

SECTION – B

Questions 6 to 12 carry 2 marks each.

- If each side of a cube is 10 cm. Find its surface area and lateral surface area.

OR

The diameters of two cones are equal. If their slant heights are in the ratio 7 : 4, find the ratio of their curved surface area.

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g}) \quad \Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \quad \Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g}) \quad \Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction $\text{C}_2\text{H}_4(\text{g}) + 2 \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g}) + \text{O}_2(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l}) \quad \Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = 285 \text{ kJ/mole}$

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web thus for reaction 9 2 in which 1 mole of H_2O and 1 mole of OH^- combine to form 2 moles of H_2O h 55 8 kj as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions 1 $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 2 $\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ 3

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