

Unit 2: Chemistry:

Kinetic Molecular Theory: pg. 135-138

The kinetic molecular theory states that:

- ❖ All matter is made of particles
- ❖ There are spaces between the particles
- ❖ All particles are moving

1. Which of the following states of matter is described in each situation below?

- a) Particles have small spaces between them and vibrate: solid
- b) Particles can slip and slide past each other resulting in no fixed shape: liquid
- c) Particles move quickly, have large spaces between them and take the volume of their container.
gas

Changes of State: pg. 139-141

2. When heat is added to matter, kinetic energy increases and thermal expansion occurs. If there is enough of an increase in energy state changes occur. What is the term used to describe each of the following state changes.

- a) solid $\rightarrow$ liquid: melting
- b) liquid $\rightarrow$ gas: evaporation
- c) gas $\rightarrow$ liquid: condensation
- d) liquid $\rightarrow$ solid: solidification or freezing
- e) solid $\rightarrow$ gas: sublimation
- f) gas $\rightarrow$ solid: deposition

Density: pg. 114 - 115

3. Use the following formula to solve the density problems below. $D = m/V$ (D=density, m=mass, v=volume) You will have this formula for your final exam!

- a) What is the density of a 2cm^3 sugar cube that has a mass of 3.18g ?

$$3.18\text{g} \div 2\text{cm}^3 = 1.59\text{g/cm}^3$$

- b) A 3mL sample of oil has a mass of 2.65g . What is the density of the oil?

$$2.65\text{g} \div 3\text{mL} = 0.88\text{g/mL}$$

Rise Of The Superbugs Answer Key

Daniela Niemeyer



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