

Question	Expected Answers	Marks	Additional Guidance
3 a	<p>measured pH > 1 OR $[H^+] < 0.1 \text{ (mol dm}^{-3}\text{)}$ ✓</p> <p>$[H^+] = 10^{-pH}$ ✓</p> <p>$K_a = \frac{[H^+][CH_3CH_2COO^-]}{[CH_3CH_2COOH]}$ OR $\frac{[H^+]^2}{[CH_3CH_2COOH]}$ ✓</p> <p>Calculate K_a from $\frac{[H^+]^2}{0.100}$ ✓</p>	4	<p>ALLOW C_2H_5 throughout question</p> <p>ALLOW $[H^+] < [CH_3CH_2COOH]$ OR $[H^+] < [HA]$ ALLOW measured pH is higher than expected ALLOW measured pH is not as acidic as expected ALLOW a quoted pH value or range > 1 and < 7 OR between 1 and 7</p> <p>ALLOW $[H^+] = \text{antilog } -pH$ OR $[H^+] = \text{inverse log } -pH$</p> <p>ALLOW $\frac{[H^+][A^-]}{[HA]}$ OR $\frac{[H^+]^2}{[HA]}$</p> <p>IF K_a is NOT given and $K_a = \frac{[H^+]^2}{0.100}$ is shown, award mark for K_a also (I.e. $K_a = \frac{[H^+]^2}{0.100}$ is automatically awarded the last 2 marks)</p>
b	<p>Marks are for correctly calculated values. Working shows how values have been derived.</p> <p>$[H^+] = 10^{-13.46} = 3.47 \times 10^{-14} \text{ (mol dm}^{-3}\text{)}$ ✓</p> <p>$[OH^-] = \frac{1.0 \times 10^{-14}}{3.47 \times 10^{-14}} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓</p>	2	<p>ALLOW $3.467368505 \times 10^{-14}$ and correct rounding to 3.5×10^{-14}</p> <p>ALLOW 0.28840315 and correct rounding to 0.29, i.e. ALLOW 0.288</p> <p>ALLOW alternative approach using pOH:</p> <p>pOH = 14 – 13.46 = 0.54 ✓ $[OH^-] = 10^{-0.54} = 0.29 \text{ (mol dm}^{-3}\text{)}$ ✓</p> <p>Correct answer gets BOTH marks</p>

Ocr F325 June 2010 Question Paper

Annelies Wilder-Smith



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