

# F321 Revision of definitions to be learnt for exam

<b>Isotopes</b>	Atoms of the same element with different numbers of neutrons and different masses.
<b>Relative atomic mass, <math>A_r</math></b>	The weighted mean mass of an atom of an element compared with one-twelfth of the mass of an atom of $^{12}\text{C}$ .
<b>Relative isotopic mass</b>	The mass of an atom of an isotope compared with one-twelfth of the mass of an atom of carbon-12.
<b>Relative formula mass</b>	The relative formula mass of a compound is equal to the sum of the individual relative atomic masses.
<b>Relative molecular mass, <math>M_r</math></b>	The weighted mean mass of a molecule compared with one-twelfth of the mass of an atom of carbon-12.
<b>Mole</b>	The amount of any substance containing as many particles as there are carbon atoms in exactly 12 g of the carbon-12 isotope.
<b>Molecular formula</b>	The number of atoms of each element in a molecule.
<b>Empirical formula</b>	The simplest whole-number ratio of atoms of each element present in a compound.
<b>Acid</b>	A species that is a proton donor.
<b>Base</b>	A species that is a proton acceptor.
<b>Concentrated Solution</b>	Solutions containing a higher number of moles of solute in a given volume of solution.
<b>Dilute Solution</b>	Solutions containing a fewer number of moles of solute in a given volume of solution.
<b>Hydrated</b>	Salt containing water molecules known as "water of crystallisation".
<b>Anhydrous</b>	A substance that contains no water of crystallisation.
<b>Oxidation</b>	Loss of electrons, or an increase in oxidation number.
<b>Reduction</b>	Gain of electrons, or a decrease in oxidation number.
<b>Redox reaction</b>	A reaction in which both reduction and oxidation take place.
<b>(First) ionisation energy</b>	The energy required to remove one electron from each atom in one mole of gaseous atoms to form one mole of gaseous $1+$ ions.
<b>(Second) ionisation energy</b>	The energy required to remove one electron from each ion in one mole of gaseous $1+$ ions to form one mole of gaseous $2+$ ions.

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**JG Myers**



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