

- (b) A solution is prepared by dissolving 10.0 g of the compound with CH_3COOH (5.00 M), and enough water (50.0 mL) to make 100 mL.

- (c) The solution acts as a buffer. Calculate the pH of the solution.

The first two values are given. 10.0 g of the compound is dissolved in 100 mL of CH_3COOH (5.00 M).

In our system, we should provide both of the species involved that would be used in calculating appropriate quantities and answer the question.

Let the compound be CH_3COOH and CH_3COO^- .

$$10.0 \text{ g} = 10.0 \text{ g} \times 10^{-3} \text{ kg} = \underline{0.01 \text{ kg}} \times 10^{-3} \text{ kg}$$

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{[\text{CH}_3\text{COO}^-]}{[\text{CH}_3\text{COOH}]} \cdot [\text{H}^+] = \frac{[\text{CH}_3\text{COO}^-]}{[\text{CH}_3\text{COOH}]} \cdot \frac{1.8 \times 10^{-5}}{1.0 \times 10^{-3}} = \underline{1.8 \times 10^{-5} \text{ mol/L}}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

$$\text{pH} = 4.74$$

Reaction: $\text{CH}_3\text{COOH} + \text{H}_2\text{O} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$
 Initial: 10.0 g, 0.0 mol, 0.0 mol, 0.0 mol
 Change: -10.0 g, +0.0 mol, +0.0 mol, +0.0 mol
 Equilibrium: 0.0 g, 0.0 mol, 0.0 mol, 0.0 mol
 The solution is homogeneous.

- (b) The solution is prepared by dissolving 10.0 g of the compound with CH_3COOH (5.00 M), and enough water (50.0 mL) to make 100 mL.

Calculate the pH of the solution.



- (c) The solution acts as a buffer. Calculate the pH of the solution.

The first two values are given. 10.0 g of the compound is dissolved in 100 mL of CH_3COOH (5.00 M).

- (d) Calculate the pH of the solution.

$$K_a = \frac{[\text{CH}_3\text{COO}^-][\text{H}^+]}{[\text{CH}_3\text{COOH}]} = \frac{[\text{CH}_3\text{COO}^-]}{[\text{CH}_3\text{COOH}]} \cdot [\text{H}^+] = \frac{[\text{CH}_3\text{COO}^-]}{[\text{CH}_3\text{COOH}]} \cdot \frac{1.8 \times 10^{-5}}{1.0 \times 10^{-3}} = \underline{1.8 \times 10^{-5} \text{ mol/L}}$$

$$\text{pH} = -\log[\text{H}^+] = -\log(1.8 \times 10^{-5}) = 4.74$$

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Ocr Chemistry A Past Papers 23

M Walker



Ocr Chemistry A Past Papers 23:

Ocr Chemistry A Past Papers 23 Book Review: Unveiling the Magic of Language

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