














Molecular Geometry Chart

# of Electron Groups	Number of Lone Pairs	Electron Pair Arrangement	Molecular Geometry	Approximate Bond Angles
2	0	linear		180°
3	0	trigonal planar		120°
	1	bent		$<120^\circ$
4	0	tetrahedral		109.5°
	1	trigonal pyramid		$<109.5^\circ$ ($\sim 107^\circ$)
	2	bent		$<109.5^\circ$ ($\sim 105^\circ$)
5	0	trigonal bipyramidal		$90^\circ, 120^\circ$
	1	see-saw		$<90^\circ, <120^\circ$
	2	T-structure		$<90^\circ$
	3	linear		180°
6	0	octahedral		$90^\circ, 90^\circ$
	1	square pyramidal		$90^\circ, <90^\circ$
	2	square planar		90°

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The Shape and Structure of Molecules Charles Alfred Coulson, 1982 **Molecular Geometry** Alison Rodger, Mark Rodger, 2014-05-16 Molecular Geometry discusses topics relevant to the arrangement of atoms. The book is comprised of seven chapters that tackle several areas of molecular geometry. Chapter 1 reviews the definition and determination of molecular geometry while Chapter 2 discusses the unified view of stereochemistry and stereochemical changes. Chapter 3 covers the geometry of molecules of second row atoms and Chapter 4 deals with the main group elements beyond the second row. The book also talks about the complexes of transition metals and f block elements and then covers the organometallic compounds and transition metal clusters. The last chapter tackles the consequences of small local variations in geometry. The text will be of great use to chemists who primarily deal with the properties of molecules and atoms. **Shape in Chemistry**

Paul G. Mezey, 1993 **Molecules and Models** Arne Haaland, 2008-03-06 This book provides a systematic description of the molecular structures and bonding in simple compounds of the main group elements with particular emphasis on bond distances, bond energies and coordination geometries. The description includes the structures of hydrogen halogen and methyl derivatives of the elements in each group. Some of these molecules are ionic, some polar covalent. The survey of molecules whose structures conform to well established trends is followed by representative examples of molecules that do not conform. We also describe electron donor acceptor and hydrogen bonded complexes. Chemists use models to systematize our knowledge, to memorize information and to predict the structures of compounds that have not yet been studied. The book provides a lucid discussion of a number of models such as the Lewis electron pair bond and the VSEPR models, the spherical and polarizable ion models and molecular orbital calculations and it outlines the successes and failures of each. **The**

Structure of Small Molecules W. J. Orville-Thomas, 1966 *Molecules* Dale Atkins, 1987-08-15 Explains how diagrams are used to represent chemical bonds and describes the structure and characteristics of molecules encountered in everyday life.

Molecules Theodore W. Gray, 2014 In his highly anticipated sequel to *The Elements* Theodore Gray demonstrates how the elements of the periodic table combine to form the molecules that make up our world. Everything physical is made up of the elements and the infinite variety of molecules they form when they combine with each other. In *Molecules* Theodore Gray takes the next step in the grand story that began with the periodic table in his best selling book *The Elements: A Visual Exploration of Every Known Atom in the Universe*. Here he explores through fascinating stories and trademark stunning photography the most interesting, essential, useful and beautiful of the millions of chemical structures that make up every material in the world. Gray begins with an explanation of how atoms bond to form molecules and compounds as well as the difference between organic and inorganic chemistry. He then goes on to explore the vast array of materials molecules can create including soaps and solvents, gums and oils, rocks and ores, ropes and fibers, painkillers and dangerous drugs, sweeteners, perfumes and stink bombs, colors and pigments and controversial compounds including asbestos, CFCs and

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