

has the K shown. If a reaction mixture has the concentrations as shown, in what direction is the reaction proceeding? Concentrations:  $[\text{COF}_2] = 1.22 \times 10^{-2} \text{ M}$

$$[\text{CO}_2] = 2.00 \times 10^{-4} \text{ M} \quad [\text{CF}_4] = 3.00 \times 10^{-4} \text{ M}$$

- A) Right to left      B) Left to Right  
C) It is at equilibrium      D) Not enough information to determine

8. The  $K_{eq}$  for the equilibrium below is 0.112 at 700.0 °C.



What is the value of  $K_{eq}$  at this temperature for the following reaction?



- A) 0.224      B) 0.0425      C) 0.112      D) 6.93      E) -0.112

9. At 200 °C, the equilibrium constant ( $K_p$ ) for the reaction  $2\text{HCl}_{(g)} \rightleftharpoons \text{H}_{2(g)} + \text{Cl}_{2(g)}$  is  $2.4 \times 10^{-4}$ . A closed vessel is charged with 26.1 atm of HCl. At equilibrium, the partial pressure of  $\text{Cl}_2$  is \_\_\_\_\_ atm.

- A) 294      B) 35.7      C) 17.9      D) 6.00      E)  $100 \times 10^{-4}$

10. In the coal-gasification process, carbon monoxide is converted to carbon dioxide via the following reaction:  $\text{CO}_{(g)} + \text{H}_2\text{O}_{(g)} \rightleftharpoons \text{CO}_{2(g)} + \text{H}_{2(g)}$ . In an experiment, 0.35 mol of CO and 0.40 mol of  $\text{H}_2\text{O}$  were placed in a 1.00-L reaction vessel. At equilibrium, there were 0.19 mol of CO remaining.  $K_{eq}$  at the temperature of the experiment is \_\_\_\_\_.

- A) 5.47      B) 0.75      C) 1.78      D) 0.56      E) 1.0

Short Answer:

1. Nitrosyl bromide decomposes according to the following equation.



A sample of NOBr (0.44 mol) was placed in a 2.00-L flask containing no NO or  $\text{Br}_2$ . If the  $K_{eq} = 58.2$ , how many moles of NO and  $\text{Br}_2$ , respectively, are in the flask at equilibrium? (4T, 1C)

2. The reaction:  $2\text{H}_2\text{O}(g) + 3\text{O}_2(g) \rightleftharpoons 2\text{H}_2\text{O}_2(g)$   $K = 3.9 \times 10^4$

has the K shown. If a reaction mixture has the concentrations as shown,

# Review Chemical Equilibrium Modern Chemistry Answers

**K Payea**



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