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# Respironics Simplygo Manual

**Marcel A. Müller**



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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho}$  32 g mol<sup>-1</sup> 790 g l<sup>-1</sup> 0 0405 l mol<sup>-1</sup> nonnumber

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web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contain 6 02 10<sup>23</sup> molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6 10<sup>23</sup> particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6 10<sup>23</sup> but as different particles have different mass their molar mass varies

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web mar 12 2018 you can define the concentration of a gas though and that calculation would be exactly the one you ve described placing x moles of a gas in a container of volume v yields a concentration  $c = \frac{x}{v}$  with units of e g pu mol over dm<sup>3</sup>

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web the molar volume has the si unit of cubic metres per mole m<sup>3</sup> mol although it is more typical to use the units cubic decimetres per mole dm<sup>3</sup> mol for gases and cubic centimetres per mole cm<sup>3</sup> mol for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22 4 l

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web jun 8 2023  $V = 1.00 \text{ mol} \times 0.08206 \text{ latm mol}^{-1} \text{K}^{-1} \times 273 \text{ K} = 22.4 \text{ l}$  in other words 1 mole of a gas will occupy 22 4 l at stp assuming ideal gas behavior at stp the volume of a gas is only dependent on number of moles of

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