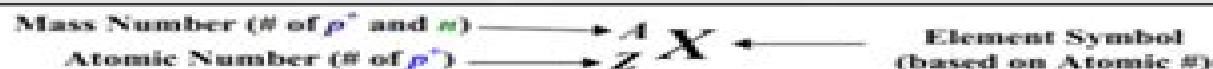


Chapter 21: Nuclear Chemistry21.1: The Nature of Nuclear Reactions

Nucleons: - the particles that make up a nucleus of an atom (protons, (${}^1_1p^+$ or 1_1H) and neutrons, (1_0n)).

Isotopes: - atoms that have different mass number but the same atomic number or number of protons.

Nuclide: a particular atom or isotope containing specific numbers of protons and neutrons



Radioactivity: - the particles and/or electromagnetic radiation that are emitted due to unstable nuclei.
 - all elements having atomic number 84 (Polonium) and greater are radioactive.

Nuclear Transmutation: - a process where radioactivity is resulted from the bombardment of nuclei by neutrons, protons or other nuclei.
 - in most cases, heavier elements are synthesized from lighter elements.

History of Radioactivity

- In 1896, **Wilhelm Roentgen** discovered X-ray by examining ray emitting from the outside of the cathode ray glass tube. It has the capability of passing through solid materials, but can be blocked by denser matters. It can also be exposed to photographic plate, resulting from an image of "seeing through" an container made of less dense material.
- In the same year, **Antoine Becquerel** discovered that uranium emits a ray onto a photographic plate in the absence of sunlight or other forms of energy.
- A few years later, **Marie and Pierre Curie** demonstrated that radiation can be emitted by other elements. They discovered two new elements (polonium and radium) based on their tendency to emit radiation (**radioactivity**).

Types of Radioactive Particle and Decay

- Alpha Particle (α particle):** - basically a helium nucleus (4_2He), commonly found during radioactive decay from heavier nuclide (the net result is to increase the neutron to proton ratio – more explanation in the next section).



- Beta Particle (β particle):** - basically an electron (${}^0_{-1}e$ or ${}^0_{-1}\beta$) that is emitted when the neutron to proton ratio is higher than the zone of stability (**a neutron is transformed to a proton and an electron as a result** – more explanation in the next section).
 - electrons have a mass number of 0 and an atomic number assignment of -1, due to its charge.



Review Nuclear Chemistry Modern Chemistry Answers

Morris Hein,Susan Arena,Cary Willard



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Modern Chemistry Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, 2001 Reviews of Modern Quantum Chemistry K. D. Sen, 2002 This important book collects together state of the art reviews of diverse topics covering almost all the major areas of modern quantum chemistry The current focus in the discipline of chemistry synthesis structure reactivity and dynamics is mainly on control A variety of essential computational tools at the disposal of chemists have emerged from recent studies in quantum chemistry The acceptance and application of these tools in the interfacial disciplines of the life and physical sciences continue to grow The new era of modern quantum chemistry throws up promising potentialities for further research Reviews of Modern Quantum Chemistry is a joint endeavor in which renowned scientists from leading universities and research laboratories spanning 22 countries present 59 in depth reviews Along with a personal introduction written by Professor Walter Kohn Nobel laureate Chemistry 1998 the articles celebrate the scientific contributions of Professor Robert G Parr on the occasion of his 80th birthday List of Contributors W Kohn M Levy R Pariser B R Judd E Lo B N Plakhutin A Savin P Politzer P Lane J S Murray A J Thakkar S R Gadre R F Nalewajski K Jug M Randic G Del Re U Kaldor E Eliav A Landau M Ehara M Ishida K Toyota H Nakatsuji G Maroulis A M Mebel S Mahapatra R Carb Dorca Nagy I A Howard N H March S B Liu R G Pearson N Watanabe S Ten no S Iwata Y Udagawa E Valderrama X Fradera I Silanes J M Ugalde R J Boyd E V Lude a V V Karasiev L Massa T Tsuneda K Hirao J M Tao J P Perdew O V Gritsenko M Gr ning E J Baerends F Aparicio J Garza A Cedillo M Galv n R Vargas E Engel A H ck R N Schmid R M Dreizler J Poater M Sol M Duran J Robles X Fradera P K Chattaraj A Poddar B Maiti A Cedillo S Guti rrez Oliva P Jaque A Toro Labb H Chermette P Boulet S Portmann P Fuentealba R Contreras P Geerlings F De Proft R Balawender D P Chong A Vela G Merino F Kootstra P L de Boeij R van Leeuwen J G Snijders N T Maitra K Burke H Appel E K U Gross M K Harbola H F Hameka C A Daul I Ciofini A Bencini S K Ghosh A Tachibana J M Cabrera Trujillo F Tenorio O Mayorga M Cases V Kumar Y Kawazoe A M K ster P Calaminici Z G mez U Reveles J A Alonso L M Molina M J L pez F Dugue A Ma anes C A Fahlstrom J A Nichols D A Dixon P A Derosa A G Zacarias J M Seminario D G Kanhere A Vichare S A Blundell Z Y Lu H Y Liu M Elstner W T Yang J Mu oz X Fradera M Orozco F J Luque P Tarakeshwar H M Lee K S Kim M Valiev E J Bylaska A Gramada J H Weare J Brickmann M Keil T E Exner M Hoffmann J Rychlewski AP Chemistry For Dummies Peter J. Mikulecky, Michelle Rose Gilman, Kate Brutlag, 2008-11-13 A practical and hands on guide for learning the practical science of AP chemistry and preparing for the AP chem exam Gearing up for the AP Chemistry exam AP Chemistry For Dummies is packed with all the resources and help you need to do your very best Focused on the chemistry concepts and problems the College Board wants you to know this AP Chemistry study guide gives you winning test taking tips multiple choice strategies and topic guidelines as well as great advice on optimizing your study time and hitting the top of your game on test day This user friendly guide helps you prepare without perspiration by developing a pre test plan organizing your study time and getting the most out of your AP course You ll get help understanding atomic

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1. Understanding the eBook Review Nuclear Chemistry Modern Chemistry Answers
 - The Rise of Digital Reading Review Nuclear Chemistry Modern Chemistry Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Review Nuclear Chemistry Modern Chemistry Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in a Review Nuclear Chemistry Modern Chemistry Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Review Nuclear Chemistry Modern Chemistry Answers
 - Personalized Recommendations

- Review Nuclear Chemistry Modern Chemistry Answers User Reviews and Ratings
- Review Nuclear Chemistry Modern Chemistry Answers and Bestseller Lists
- 5. Accessing Review Nuclear Chemistry Modern Chemistry Answers Free and Paid eBooks
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 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Review Nuclear Chemistry Modern Chemistry Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Review Nuclear Chemistry Modern Chemistry Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Review Nuclear Chemistry Modern Chemistry Answers
 - Setting Reading Goals Review Nuclear Chemistry Modern Chemistry Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Review Nuclear Chemistry Modern Chemistry Answers
 - Fact-Checking eBook Content of Review Nuclear Chemistry Modern Chemistry Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
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