

Name: _____

Your test is on Friday, 12/18/15

Chemical Reactions Study GuideBalancing Chemical Equations

The first step to balancing chemical reaction equations is to be able to accurately count the number of each type of element on both the reactants and products side of the arrow. This is done by multiplying the **coefficient** (big number in front of an element symbol) with the **subscript** (little number found after the element symbol). **Remember:** if no coefficient or subscript is written, then there is an implied, imaginary "1".

Examples: 3Br_2 $\leftarrow$ there are 6 bromine atoms ($3 \times 2 = 6$)

CaO $\leftarrow$ there is 1 calcium and 1 oxygen (in each case $1 \times 1 = 1$). The "1"'s are implied.

Practice: How many of each type of atom is represented.

3Cl_2 6 Fe 1 2Na 16 MgF_2 1 Mg; 2F

I_2 2 4Ni 4 $2 \text{Li}_2\text{O}$ 4 Li; 2 O H_2O_2 2 H; 2O

H_2SO_4 2 H; 1 SO₄ $2 \text{Pb}(\text{NO}_3)_2$ 2 Pb; 4 NO₃

Now that we've learned to count, it's time to put our skills into action by altering the numbers of certain elements to suit our desires. In any case, our desires are to have the same number of each type of element on both the **products** and **reactants** side of the equation. We do this by **changing the coefficients**.

Practice: Balance each of the following chemical reaction equations. Check your answer by counting each element on each side when you think you're done.



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