Chapter 4 - Practice Problems Solutions

 Complete and balance the following reactions. Remember to form the stable product and then balance. Assume the product is solid. What type of reactions are these?

____ Mg (s) + ____ O₂ (g)
$$\rightarrow$$
 ____ Na (s) + ____ Cl₂ (g) \rightarrow 2 Mg (s) + O₂ (g) \rightarrow 2 MgO (s) 2 Na (s) + Cl₃ (g) \rightarrow 2 NaCl (s)

These are all combination reactions.

2) Balance the following reactions. What type of reactions are these?

$$C_3H_8(g) + __O_2(g) \rightarrow _CO_2(g) + __H_2O(g)$$

 $C_3H_8(g) + 5 O_2(g) \rightarrow _CO_2(g) + 4 H_2O(g)$
 $C_6H_6(g) + _O_2(g) \rightarrow _CO_2(g) + _H_2O(g)$
 $C_6H_6(g) + 15 O_2(g) \rightarrow _CO_2(g) + 6 H_2O(g)$

These are both combustion reactions.

3) Balance the following reactions. What type of reactions are these?

$$H_2O(I)$$
 → _____ $H_2(g)$ + _____ $O_2(g)$ _____ $KCIO_3(s)$ → _____ $KCI(s)$ + _____ $O_2(g)$
2 $H_2O(I)$ → 2 $H_2(g)$ + $O_2(g)$ 2 $KCIO_3(s)$ → 2 $KCI(s)$ + 3 $O_2(g)$

These are both decomposition reactions.

4) Balance the following reactions. What type of reactions are these?

____ Mg (s) + ____ HCl (aq)
$$\rightarrow$$
 ____ MgCl₂ (aq) + ____ H₂ (g)
Mg (s) + 2 HCl (aq) \rightarrow MgCl₂ (aq) + H₂ (g)
____ Ag (s) + ____ H₂S (aq) \rightarrow ____ Ag₂S (s) + ____ H₂ (g)
2 Ag (s) + H₃S (aq) \rightarrow Ag₂S (s) + H₂ (g)

These are both single displacement reactions.

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