

1. There was no organization in the instructions. The instructions should be written in a "...three-part structure of instruction, namely, first operation, second skill." The instructions presented above listed all things with step one, there's no instruction on how to turn on the device. There's no review of instruction or how to implement presented in the instructions.
  2. There is no defined audience for these instructions. With a clear audience, that will depict the "...the level of familiarity your readers have with the topic presented." If the target audience is for people who are familiar with technology, the instructions don't have to go as in-depth detail as if the target audience was asked for people that are not as tech-savvy. Then the instructions would have to be step by step.
  3. There is no equipment needed list. That should be in the beginning of the instructions manual. With the equipment list, it could leave the reader with unanswered questions. What cord attaches to the computer to upload the recording? Questions such as that could be easily answered with an equipment list.
- Moreno, Caroline, et al. "P.T. Writing Instructions." *Technical Writing Handbook*, University of Illinois. 2 Jan. 2008. [www.english.illinois.edu/technicalwriting/pt.html](http://www.english.illinois.edu/technicalwriting/pt.html).  
"CCRM 2008: THE FINAL REPORT." 2008.  
http://www.ccrp.org/initial2008/initial2008finalreport/initial2008.htm.

# Sample Poorly Written Manual

**Heather Graves,Roger Graves**



## **Sample Poorly Written Manual:**

## Decoding **Sample Poorly Written Manual**: Revealing the Captivating Potential of Verbal Expression

In a period characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its power to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Sample Poorly Written Manual**," a mesmerizing literary creation penned by a celebrated wordsmith, readers attempt an enlightening odyssey, unraveling the intricate significance of language and its enduring affect our lives. In this appraisal, we shall explore the book is central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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web the molar volume of a gas expresses the volume occupied by 1 mole of that respective gas under certain temperature and pressure conditions the most common example is the molar volume of a gas at stp standard temperature and pressure which is equal to  $22.4 \text{ l}$  for 1 mole of any ideal gas at a temperature equal to  $273.15 \text{ k}$  and a pressure

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web the molar volume will be the volume occupied by one molar mass 32 g of the liquid expressing the density in liters instead of ml we have  $v = \frac{m}{\rho}$  32 g mol<sup>-1</sup> 790 g l<sup>-1</sup> 0 0405 l mol<sup>-1</sup> nonnumber

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web the molar volume of a gas from avogadro s law the volume is directly proportional to the number of moles the important point to keep in mind is that n can be moles of any gas since one mole always contain 6 02 10<sup>23</sup> molecules and one molecule of any gas at the same temperature will have the same kinetic energy based on this principle it is found

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web apr 22 2020 molar volume of gases 1 counting in moles one mole is defined as 6 10<sup>23</sup> particles this is the avogadro number an unchanging constant whether you are counting hydrogen atoms carbon dioxide molecules or carbonate ions a mole is always 6 10<sup>23</sup> but as different particles have different mass their molar mass varies

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web the molar volume has the si unit of cubic metres per mole m<sup>3</sup> mol although it is more typical to use the units cubic decimetres per mole dm<sup>3</sup> mol for gases and cubic centimetres per mole cm<sup>3</sup> mol for liquids and solids

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web zero degrees celsius or 273 k is defined as standard temperature and one atmosphere or 760 mm hg is defined as standard pressure the volume occupied by one mole of a gas is called molar volume therefore the molar volume of a gas at standard temperature and pressure stp is 22 4 l

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