

Equilibrium and Reaction Rates

(answers at the end)

Q: What is Le Chatelier's law (Le Chatelier's principle)?

A: Le Chatelier's principle states that:

When a system at equilibrium is subjected to a change in conditions,

the system will adjust itself to counteract the change and restore equilibrium.

Q: What is the effect of increasing the concentration of a reactant on the equilibrium position?

A: Increasing the concentration of a reactant will shift the equilibrium position to the right, favoring the formation of products.

Q: What is the effect of increasing the temperature on the equilibrium position?

A: Increasing the temperature will shift the equilibrium position to the right, favoring the formation of products.

Q: What is the effect of increasing the pressure on the equilibrium position?

A: Increasing the pressure will shift the equilibrium position to the left, favoring the formation of reactants.

Q: What is the effect of increasing the volume on the equilibrium position?

A: Increasing the volume will shift the equilibrium position to the right, favoring the formation of products.

Q: What is the effect of increasing the catalyst on the equilibrium position?

A: Increasing the catalyst will not shift the equilibrium position, but it will increase the rate of reaction.

- the concentration of reactants
- the concentration of products
- the concentration of catalyst
- the temperature of the system
- the pressure of the system

Q: What is the effect of increasing the volume on the equilibrium position?

A: Increasing the volume will shift the equilibrium position to the right, favoring the formation of products.

Reactions Rates And Equilibrium Section Review Answers

Daniela Niemeyer



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