Worksheet- Reaction Rates

Use this reaction for the questions below: $C_0H_{22}O_0$ (s) + 6 $O_2(g) \rightarrow$ 6 H_2O (g) + 6 CO_2 (g)

- 1. What happens to the concentrations of:
 - a. C₆H₁₂O₆ & O₂ as the reaction proceeds →?
- b. H₂O + CO₂ as the reaction proceeds →?
- 2. According to the collision theory, what 3 circumstances are needed for CoHizOo & Oz to react?
- 3. What is the activation energy for a chemical reaction?
- 4. Use the equation & the collision theory to explain:

$$C_6H_{12}O_6(s) + 6 O_2(g) \rightarrow 6 H_2O(g) + 6 CO_2(g)$$

	Change in condition:	Does this <u>increase</u> or <u>decrease</u> the rate of reaction?	Explain why
a.	Increasing the temperature	Exi (normanca (apenda up)	Ext Molecules more faster to collide more - Increased on rate
ь.	Increasing the concentration of CoHorOs		
C.	Decreasing the concentration of O ₂		
d.	Increase the surface area by chewing up food in your mouth		
6.	Decreasing the temperature		
f.	Increasing the pressure in the container		
9	Decreasing the concentration of H ₂ O		
h.	Increasing the volume of the container the reaction occurs in		
	Increasing the concentration of CO ₂		
1	Using a catalyst (like solivary amylase)		

- On the accompanying energy diagram, <u>label</u> the following terms:
 - a. reactants b. products c. activation energy
- On the graph to the right, <u>draw</u> and <u>label</u> what this diagram would look like if a <u>catalyst</u> was added to the reaction.
- 7. Graph reading
 - a. How much energy (#) do the reactants have?
 - b. How much energy (#) do the products have?
 - c. How much energy (#) is required to activate this un-catalyzed reaction?
 - d. Is this reaction endothermic or exothermic? How do you know?



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