

Identifying Reaction Types & Balancing Equations

Step 1: Name the type of chemical reaction	Step 2: Balance the following chemical equations:
Ex. Combination Reaction	$2 \text{ N}_2 + 6 \text{ H}_2 \rightarrow 4 \text{ NH}_3$
1.	$\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$
2.	$\text{NaCl} + \text{F}_2 \rightarrow \text{NaF} + \text{Cl}_2$
3.	$\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$
4.	$\text{AgNO}_3 + \text{MgCl}_2 \rightarrow \text{AgCl} + \text{Mg(NO}_3)_2$
5.	$\text{AlBr}_3 + \text{K}_2\text{SO}_4 \rightarrow \text{KBr} + \text{Al}_2(\text{SO}_4)_3$
6.	$\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
7.	$\text{C}_3\text{H}_8 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
8.	$\text{FeCl}_3 + \text{NaOH} \rightarrow \text{Fe(OH)}_3 + \text{NaCl}$
9.	$\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$
10.	$\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
11.	$\text{Ag}_2\text{O} \rightarrow \text{Ag} + \text{O}_2$
12.	$\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2$
13.	$\text{K} + \text{MgBr}_2 \rightarrow \text{KBr} + \text{Mg}$
14.	$\text{HNO}_3(\text{aq}) + \text{Ba(OH)}_2(\text{aq}) \rightarrow \text{Ba(NO}_3)_2(\text{aq}) + \text{H}_2\text{O (l)}$
15.	$\text{C}_3\text{H}_{12}(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + \text{H}_2\text{O(g)}$
16.	$\text{Al(s)} + \text{Fe}_2\text{O}_3(\text{aq}) \rightarrow \text{Al}_2\text{O}_3(\text{aq}) + \text{Fe(s)}$
17.	$\text{Al(s)} + \text{O}_2(\text{g}) \rightarrow \text{Al}_2\text{O}_3(\text{s})$
18.	$\text{P}_4 + \text{O}_2 \rightarrow 2 \text{ P}_2\text{O}_5$
19.	$2 \text{ CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
20.	$\text{H}_2 + \text{O}_2 \rightarrow 2 \text{ H}_2\text{O}$

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