

Name \_\_\_\_\_ Date \_\_\_\_\_ Class \_\_\_\_\_

**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

## Solutions

### Section 15.1 What are solutions?

In your textbook, read about the characteristics of solutions.

Use each of the terms below just once to complete the passage.

|            |          |         |          |
|------------|----------|---------|----------|
| immiscible | liquid   | soluble | solution |
| insoluble  | miscible | solvent | solvent  |

Air is air! (1) solution of oxygen gas dissolved in nitrogen gas. The oxygen is air is the (2) solvent and nitrogen is the (3) solvent. Because oxygen gas dissolves in a solvent, oxygen gas is air is (4) soluble substances. A substance that does not dissolve is (5) insoluble. (6) liquid solutions are the most common type of solutions. If one liquid is soluble in another liquid, such as acetic acid in water, the two liquids are (7) miscible. However, if one liquid is insoluble in another, the liquids are (8) immiscible.

Read about solvation in aqueous solutions in your textbook.

The diagram shows the hydration of solid sodium chloride to form an aqueous solution. Use the diagram to answer the following questions.



8. Hydration is solvation in which the solvent is water. What is solvation?

Solvation is the process of surrounding solute particles with solvent particles to form a solution.

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**CHAPTER 15** **STUDY GUIDE FOR CONTENT MASTERY**

### Section 15.1 continued

10. As sodium chloride dissolves in water, what happens to the sodium and chloride ions?

The sodium and chloride ions are separated and surrounded by the water molecules.

11. Explain the attraction of the water molecules around the sodium ions and chloride ions.

Because the sodium ion is positively charged, it attracts the negatively charged portion of the water molecule (the oxygen atom) and repels the positively charged portion of the water molecule (the hydrogen atoms). Because the chloride ion is negatively charged, it attracts the positively charged portion of the water molecule and repels the negatively charged portion.

12. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?

The attraction between the water molecules and the sodium and chloride ions is greater than the attraction between the sodium and chloride ions. The greater strength of attraction between the water molecules and the ions is what causes the solvation process to occur.

13. List three ways that the rate of solvation may be increased.

stirring or shaking the solution, breaking the solute into smaller pieces, and heating the solvent

In your textbook, read about heat of solution, solubility, and factors that affect solubility.

For each statement below, write true or false.

True \_\_\_\_\_ 14. The overall energy change that occurs when a solution forms is called the heat of solution.

False \_\_\_\_\_ 15. Solubility is a measure of the maximum amount of solute that dissolves in a given amount of solvent at a specified temperature and pressure.

False \_\_\_\_\_ 16. Solvation continues as long as the solvation rate is less than the crystallization rate.

True \_\_\_\_\_ 17. In a saturated solution, solvation and crystallization are in equilibrium.

True \_\_\_\_\_ 18. Additional solute can be dissolved in an unsaturated solution.

True \_\_\_\_\_ 19. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

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# Reaction Spontaneity Study Guide For Content Mastery

**Kay Mohanna, Ruth Chambers, David  
Wall**



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general teaching methods for K 12 classrooms that promote student understanding while emphasizing contemporary issues including differentiating your instruction creating a learning community and making modifications in instruction based on student differences This edition offers new content in every chapter on teaching English language learners in all classrooms Other unique coverage includes ways to manage lesson delivery how to apply motivational techniques for instruction and assessment and how to work with colleagues and parents All the content is aligned to professional standards The book s thorough coverage of classroom management and discipline includes a discussion of dynamic ways to create a positive learning environment

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