



# Chem!stry

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## Rate of Reaction – Assignment One – Answers

### Question 1.

- a) Write a balanced chemical equation, including state symbols, for the reaction between zinc and dilute sulfuric acid.



- b) Give **three** ways by which the speed of the reaction between zinc and hydrochloric acid can be measured.

Measure the volume of hydrogen gas produced against time.

Measure the change in pH against time.

Measure the change in temperature against time.

Measure the change in mass of the reaction against time as  $\text{H}_2\text{(g)}$  is allowed to escape.

- c) Give **three** ways in which the speed of the reaction between zinc and dilute sulphuric acid can be increased.

Increase the temperature of the reaction.

Increase the surface area (to volume ratio) of the zinc.

Increase the concentration of the sulfuric acid.

Add a catalyst to the reaction.

### Question 2.

Equal lengths of magnesium ribbon were added to  $3 \text{ mol / dm}^3$  hydrochloric acid and to  $3 \text{ mol / dm}^3$  sulfuric acid. When the reactions had stopped, some metal remained.

- a) Write the equations for the reactions.



- b) With which acid would the reaction be faster at the start? Briefly explain your answer.

Sulfuric acid,  $\text{H}_2\text{SO}_4$ . Sulfuric acid is dibasic, therefore 1 mol of sulfuric acid will ionize to produce 2 mol of hydrogen ions. Hydrochloric acid is monobasic, therefore 1 mol of hydrochloric acid will ionize to produce 1 mol of hydrogen ions. Sulfuric acid contains a greater concentration of hydrogen ions compared to the hydrochloric acid. There will be a greater frequency of effective collisions between the hydrogen ions and magnesium in the reaction using sulfuric acid.

# Reaction Rate Edgenuity Answers

**DJ Losen**

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