

Chapter 1: REACTION KINETICS

1.1 REACTION KINETICS

- a) Arrhenius equation: $k = A e^{-E_a/RT}$
- b) Transition state theory: $k = \frac{k_B T}{h} e^{-E_a/RT}$
- c) Reaction coordinate diagram: E_a is activation energy
- d) Rate law: $r = k [A]^m [B]^n$
- e) Order of reaction: $m+n$
- f) Half-life: $t_{1/2}$
- g) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- h) Rate constant: k
- i) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- j) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- k) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- l) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- m) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- n) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- o) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- p) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- q) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- r) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- s) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- t) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- u) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- v) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- w) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- x) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- y) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$
- z) Rate of reaction: $r = \frac{1}{V} \frac{d[A]}{dt}$

REACTION RATE

- a) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$
- b) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$
- c) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$
- d) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$

Reaction rate is the change in concentration of a reactant or product per unit time.

- e) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$
- f) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$
- g) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$

- h) Reaction rate: $r = \frac{1}{V} \frac{d[A]}{dt}$



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Reaction Energy And Reaction Kinetics Study Guide

KJ Lindholm-Leary



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