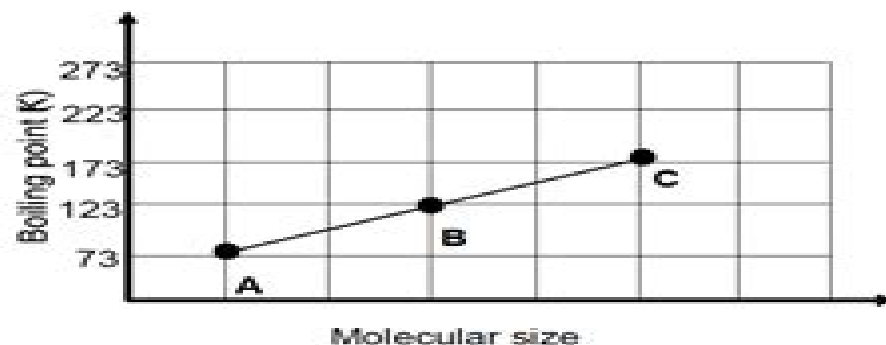


QUESTION 2

- 2.1 The graph of molecular size versus the boiling point is given below.
The letters **A**, **B** and **C** represent the compounds CH_4 , C_2H_6 and C_3H_8 respectively.



- 2.1.1 Define the term *boiling point*. (2)
- 2.1.2. Describe the trend in the boiling points of the compounds as shown by the graph. (2)
- 2.1.3. Explain the answer to QUESTION 2.1.2 by referring to MOLECULAR SIZE, TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)
- 2.1.4 Which ONE of the compounds (**A**, **B** or **C**) has the HIGHEST vapour pressure? Explain the answer by referring to the data on the graph. (2)
- 2.2 Consider the two molecules in the table below.

NAME OF SUBSTANCE	FORMULA	MOLACULAR MASS (g)	BOILING POINT ($^{\circ}\text{C}$)
Ammonia	NH_3	17	- 33
Phosphine	PH_3	x	- 87,4

- 2.2.1 Determine the molecular mass (x) of phosphine (1)
- 2.2.2 Explain the difference in the boiling points by referring to the TYPE and STRENGTH of INTERMOLECULAR FORCES. (3)

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Table of Contents Physical Science Grd11 2014 March Exam View Question Paper

1. Understanding the eBook Physical Science Grd11 2014 March Exam View Question Paper
 - The Rise of Digital Reading Physical Science Grd11 2014 March Exam View Question Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Physical Science Grd11 2014 March Exam View Question Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Physical Science Grd11 2014 March Exam View Question Paper
 - User-Friendly Interface

4. Exploring eBook Recommendations from Physical Science Grd11 2014 March Exam View Question Paper
 - Personalized Recommendations
 - Physical Science Grd11 2014 March Exam View Question Paper User Reviews and Ratings
 - Physical Science Grd11 2014 March Exam View Question Paper and Bestseller Lists
5. Accessing Physical Science Grd11 2014 March Exam View Question Paper Free and Paid eBooks
 - Physical Science Grd11 2014 March Exam View Question Paper Public Domain eBooks
 - Physical Science Grd11 2014 March Exam View Question Paper eBook Subscription Services
 - Physical Science Grd11 2014 March Exam View Question Paper Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
 - Physical Science Grd11 2014 March Exam View Question Paper Compatibility with Devices
 - Physical Science Grd11 2014 March Exam View Question Paper Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Physical Science Grd11 2014 March Exam View Question Paper
 - Highlighting and Note-Taking Physical Science Grd11 2014 March Exam View Question Paper
 - Interactive Elements Physical Science Grd11 2014 March Exam View Question Paper
8. Staying Engaged with Physical Science Grd11 2014 March Exam View Question Paper
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Physical Science Grd11 2014 March Exam View Question Paper
9. Balancing eBooks and Physical Books Physical Science Grd11 2014 March Exam View Question Paper
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Physical Science Grd11 2014 March Exam View Question Paper
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Physical Science Grd11 2014 March Exam View Question Paper
 - Setting Reading Goals Physical Science Grd11 2014 March Exam View Question Paper
 - Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Physical Science Grd11 2014 March Exam View Question Paper
 - Fact-Checking eBook Content of Physical Science Grd11 2014 March Exam View Question Paper
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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