

**EXAMINATION DATA SHEET FOR THE PHYSICAL SCIENCES
(CHEMISTRY)**

TABLE 1 PHYSICAL CONSTANTS

| NAME | SYMBOL | VALUE |
|---------------------------------|-----------|---|
| Magnitude of charge on electron | e | $1,6 \times 10^{-19} \text{ C}$ |
| Mass of an electron | m_e | $9,1 \times 10^{-31} \text{ kg}$ |
| Standard pressure | p° | $1,01 \times 10^5 \text{ Pa}$ |
| Molar gas volume at STP | V_m | $22,4 \text{ dm}^3 \cdot \text{mol}^{-1}$ |
| Standard temperature | T° | 273 K |
| Avogadro's constant | N_A | $6,02 \times 10^{23} \text{ mol}^{-1}$ |
| Faraday's constant | F | $96\,500 \text{ C} \cdot \text{mol}^{-1}$ |

TABLE 2 CHEMISTRY FORMULAE

| | | |
|---|---|---------------------|
| $n = \frac{m}{M}$ | $n = \frac{N}{N_A}$ | $n = \frac{V}{V_m}$ |
| $c = \frac{n}{V}$ OR $c = \frac{m}{MV}$ | $K_w = [\text{H}_3\text{O}^+] \cdot [\text{OH}^-] = 1 \times 10^{-14}$ at 25 °C (298 K) | |
| $q = It$ $q = nF$ | $E_{\text{cell}}^\circ = E_{\text{cathode}}^\circ - E_{\text{anode}}^\circ$ $E_{\text{cell}}^\circ = E_{\text{oxidising agent}}^\circ - E_{\text{reducing agent}}^\circ$ | |

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