

Focus 3: States of Matter

3A: The Nature of Gases

- i. Pressure is the amount of force exerted by a gas over a certain area.
 - A. $P = F/A$
 - B. The SI unit of pressure is the pascal, where 1 Pa = 1 kg/m³s²
 - C. 1 bar = 10⁵ Pa
 - D. 1 atm = 1.01325 × 10⁵ Pa
 - E. 1 atm = 760 mm/760 mm Hg

3B: The Gas Laws

- ii. **Boyle's Law** is one that takes place at a constant temperature.
 - A. An isotherm is a curve for a constant-temperature plot.
- iii. **Boyle's Law** pressure is inversely proportional to volume.
 - A. $P_1V_1 = P_2V_2$
- iv. **Avogadro's Principle** states that gas molecules occupy the same volume regardless of chemical identity, given that it is under constant pressure and temperature.
- v. The **ideal gas law** summarizes the relation between pressure, volume, and amount of a substance.
 - A. $PN = nRT$
 1. P = pressure
 2. V = volume
 3. n = moles
 4. R = ideal gas constant which is equal to:

TABLE 3B.1 The Gas Constant, R

$$8.314\text{ J K}^{-1}\text{ mol}^{-1} = 10^3\text{ J K}^{-1}\text{ mol}^{-1}$$

$$8.314\text{ J K}^{-1}\text{ mol}^{-1} = 10^3\text{ L bar K}^{-1}\text{ mol}^{-1}$$

$$8.314\text{ J K}^{-1}\text{ mol}^{-1} = 8.314\text{ L atm K}^{-1}\text{ mol}^{-1}$$

$$8.314\text{ J K}^{-1}\text{ mol}^{-1} = 8.314\text{ J K}^{-1}\text{ mol}^{-1}$$

$$8.314\text{ J K}^{-1}\text{ mol}^{-1} = 8.314\text{ L Torr K}^{-1}\text{ mol}^{-1}$$

$$R = \frac{P_1V_1}{nT_1}$$

$$\text{vi. Combined Gas Law: } \frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

vii. STP = standard temperature and pressure is 0 degrees C and 1 atm

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Mark L. Strauss



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