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## Financial Assistance Request Letter

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July 23, 2050

Drake Feeney  
Program Coordinator  
BrandKat  
Austin, TX 73301

Dear Mr. Feeney,

I trust this letter reaches you in good health. My name is **[Your Name]**, and I am writing to request financial assistance from BrandKat due to unforeseen medical expenses following a recent hospitalization.

Despite my best efforts to manage my finances, I find myself in a difficult position. I was recently diagnosed with a serious health condition that required surgery, leading to significant medical bills and lost wages during my recovery period.

I am seeking assistance for \$3,000 to help cover my medical bills and rent for the next few months. This support would greatly alleviate my financial burden and allow me to focus on my recovery without the stress of mounting expenses.

# **Sample Letter For Financial Assistance**

**Thirumalaisamy P.**  
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web question 1 find the mole fraction of methanol  $\text{CH}_3\text{OH}$  and water in a solution prepared by dissolving  $4.5 \text{ g}$  of alcohol in  $40 \text{ g}$  of  $\text{H}_2\text{O}$  molar mass of  $\text{H}_2\text{O}$  is  $18 \text{ g mol}^{-1}$  and molar mass of  $\text{CH}_3\text{OH}$  is  $32 \text{ g mol}^{-1}$  solution moles of  $\text{CH}_3\text{OH}$   $4.5 / 32 = 0.14 \text{ mol}$  moles of  $\text{H}_2\text{O}$   $40 / 18 = 2.2 \text{ mol}$  therefore according to the equation

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web we used mole fractions in gen chem 1 to describe gas mixtures where the partial pressure of a gas was its mole fraction  
time the total pressure  $p_a = x_a p_{\text{total}}$  mole fraction  $x_a = \frac{\text{moles of substance a}}{\text{total moles}}$   
a is often denoted by  $x_a$  and

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web 1 in the given mixture the sum of all the mole fractions is equal to one  $x_a + x_b + \dots = 1$  2 the multiplication of the mole fraction  
by 100 gives the mole percentage 3 mole fraction is a unitless and dimensionless expression advantages of molar fraction  
mole fraction is not dependent on the temperature

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web nov 6 2023  $x_a \text{ mol a} + x_b \text{ mol b} = 1$  note in the given mixture the sum of all the mole fractions  
present is equal to one  $x_a + x_b = 1$  when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is  
a unitless and dimensionless expression image will be uploaded soon

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ways to represent the concentration of the solution it is defined as the number of molecules of component a divided by the  
total number of molecules of all the components in the solution

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 - 30 = 70 g benzene's molar  
mass  $C_6H_6 = 78 \text{ g/mol}$  hence  $C_6H_6$ 's number of moles  $\frac{30}{78} = 0.3846 \text{ mol}$  molar mass of carbon tet  $CCl_4 = 154 \text{ g/mol}$

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g of water calculate the molarity and molality of the solution given  $n(H) = 1$   $n(C) = 12$   $n(O) = 16$  given mass of solute urea 0.211 g mass  
of solvent water 100 g 0.1 kg to find molarity of solution solution

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web solution moles of  $CH_3OH = 0.1722 \text{ mol}$  moles of  $H_2O = 0.073 \text{ mol}$  therefore according to the equation mole fraction  
of  $CH_3OH = \frac{0.1722}{0.1722 + 0.073} = 0.703$  to solve more examples on mole fraction formula and practice more  
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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0.449 molal KBr that contains 2.92 kg of water 1 molality moles solute divided by kilograms solute  $0.449 \text{ mol/kg} \times 2.92 \text{ kg} \times 131.108 \text{ mol of KBr} = 2 \text{ moles}$  times molar mass equals grams  $131.108 \text{ mol} \times 119.0023 \text{ g/mol} = 156 \text{ g KBr}$

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web calculate the mass percent mole fraction molality and molarity of hexane  $C_6H_{14}$  and  $CH_3Cl$  when it is mixed with heptane  $C_7H_{16}$  in a 35.0 : 55.0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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web jan 8 2022  $\text{mole fraction} \times 100$  the sum of all of the mole percents of a mixture equals 100 mole fraction properties and advantages the mole fraction offers advantages over some of the other units of concentration unlike molarity mole fraction is *worksheet mole fraction ms fleeming* - Jun 11 2023

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web a derivatization that achieves this goes as  $\text{molarity} \times \text{moles of solute} = \text{volume of solution in liter} \times \text{moles} = \text{molarity} \times \text{volume in liter} = \text{mass moles} \times \text{molar mass}$  numerical answers 1

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web since the mole fraction of the  $\text{O}_2$  is  $0.050$  it contributes  $0.050$  or if you will  $5.0\%$  of the total pressure thus total pressure is  $0.210 / 0.050 = 4.2 \text{ atm}$  problem 9 a sample of  $1.43 \text{ g}$  of helium and an unweighed quantity of  $\text{O}_2$  are mixed in a flask at room temperature

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