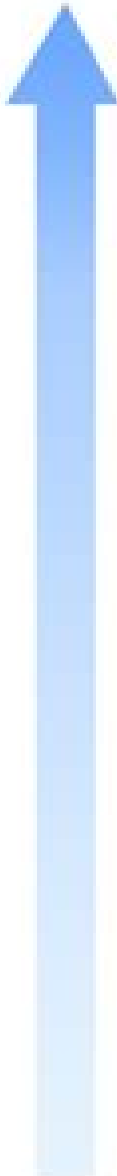


Reactivity Series

Metal (Ion)	Reactivity	Most reactive
Potassium (K ⁺)	Reacts with water	
Sodium (Na ⁺)		
Lithium (Li ⁺)		
Barium (Ba ²⁺)		
Strontium (Sr ²⁺)		
Calcium (Ca ²⁺)		
Magnesium (Mg ²⁺)	Reacts with acid	
Aluminum (Al ³⁺)		
Manganese (Mn ⁺)		
Zinc (Zn ²⁺)		
Chromium (Cr ²⁺)		
Iron (Fe ²⁺)		
Cadmium (Cd ²⁺)		
Cobalt (Co ²⁺)		
Nickle (Ni ²⁺)		
Tin (Sn ²⁺)		
Lead (Pb ²⁺)		
Hydrogen (H ⁺)	Included for comparison	
Antimony (Sb ²⁺)	Highly unreactive	
Bismuth (Bi ²⁺)		
Copper (Cu ²⁺)		
Mercury (Hg ²⁺)		
Silver (Ag ⁺)		
Gold (Au ²⁺)		
Platinum (Pt ⁺)	Least reactive	

Reactivity Series Pogil

Mike Jess

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

Reactivity Series Pogil:

Reactivity Series & Displacement Reactions, 2006 **Reactivity Series**, 2006 Chemical Reactions, Materials and Particles Gill Murphy, 2006-03 Specials are a teaching resource designed for KS3 students whose literacy skills are considerably lower than their age. These books have an older format to counteract the simple text and cover various topics. They include activities, visuals, and assessment sheets as well as teacher pages. **Mechanism and Reactivity for a Series of Hydrolytic Reactions** Terry Carl Pletcher, 1969 **Material Changes and Reactions** Chris Oxlade, 2007. What happens in a redox reaction? How do materials change when they melt or freeze? Why are chemical reactions so important? This title explores how elements melt, boil, and freeze; how they combine with each other to make new substances; and how chemical reactions take place every day in factories and in your home. You will also find several experiments that can be done at home.

Chemical Reactions Simon de Pinna, 2007-01-12. Introduces chemical reactions covering how atoms, compounds, and mixtures form acids, bases, and salts, and what happens when these come in contact with each other under certain conditions.

Chemical Reactions Carol Baldwin, 2006. Describes the different types of chemical reactions and how temperature, concentration, particle size, and catalysts affect the reaction rate. **Reactivity in Organic Chemistry** G. W. Klumpp, 1982.

Solvent Effects and Chemical Reactivity Orlando Tapia, Juan Bertrán, 2003-07-31. This book presents an up-to-date view of theories, practical methods, and applications of solvent effects and chemical reactivity in condensed phases. Subjects treated include continuum solvation models, the theoretical basis for the treatment of solvent effects in density functional theory, Monte Carlo simulations of chemical reactions in solution, DFT molecular dynamics simulations crossing the transition state in solution, valence bond multi-state approach to chemical reactions in solution, quantum theory of solvent effects, and chemical reactions. The approaches taken, as well as the resulting findings, are discussed in detail, thus covering a large part of the methodology currently used in this field. Audience: This volume will be useful to graduate students in chemistry, physical chemistry, and biochemistry; to research workers with a background in quantum chemistry and quantum mechanics; to pure and applied quantum chemists; and to industrial molecular modellers. *The Reaction Path in Chemistry: Current Approaches and Perspectives* D. Heidrich, 2013-01-07. The so-called reaction path (RP) with respect to the potential energy or the Gibbs energy (free enthalpy) is one of the most fundamental concepts in chemistry. It significantly helps to display and visualize the results of the complex microscopic processes forming a chemical reaction. This concept is an implicit component of conventional transition state theory (TST). The model of the reaction path and the TST form a qualitative framework which provides chemists with a better understanding of chemical reactions and stirs their imagination. However, an exact calculation of the RP and its neighbourhood becomes important when the RP is used as a tool for a detailed exploring of reaction mechanisms and particularly when it is used as a basis for reaction rate theories above and beyond TST. The RP is a theoretical instrument that now forms the theoretical heart of direct dynamics. It is particularly useful for the interpretation

of reactions in common chemical systems A suitable definition of the RP of potential energy surfaces is necessary to ensure that the reaction theories based on it will possess sufficiently high quality Thus we have to consider three important fields of research Analysis of potential energy surfaces and the definition and best calculation of the RPs or at least of a number of selected and chemically interesting points on it The further development of concrete versions of reaction theory beyond TST which are applicable for common chemical systems using the RP concept *Chemistry & Chemical Reactivity* John C. Kotz, Keith F. Purcell, 1987 Reactivity Patterns in a Series of Novel Monothiocarbamates, Dithiocarbamates, and Sulfur-rich Dithiolates Donald M. Baird, 1982 *Chemistry and Chemical Reactivity* John C. Kotz, Keith F. Purcell, 1987

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