

# KEY

## Unit 9 Outline- Thermodynamics, Reaction Rates, & Chemical Equilibrium Chapters 17 & 18

### Essential Skills/ State Standards:

- 9-1. Describe how temperature and heat flow in terms of the motion of molecules (or atoms). (CSS: 7a)  
 9-2. Know that chemical processes & phase changes can either release energy (exothermic) or absorb energy (endothermic). (CSS: 7b, 7c)  
 9-3. Solve problems involving heat flow and temperature changes, using known values of specific heat ( $q = m \cdot c \cdot \Delta T$ ) & latent heat of phase change ( $q = m \cdot \Delta H$ ). (CSS: 7d)  
 9-4. Know the rate of reaction is the decrease in concentration of reactants or the increase in concentration of products with time & depends on such factors as: concentration, temperature, and pressure. (CSS: 8a, 8b)  
 9-5. Know the definition and role of activation energy & catalyst in a chemical reaction. (CSS: 8d\*, 8c)  
 9-6. Know how to use Le Chatelier's principle to predict the effect of changes in concentration, temperature, and pressure. (CSS: 9a)  
 9-7. Know that equilibrium is established when the forward and reverse reaction rates are equal. (CSS: 9b)

### Vocabulary

- |                |                   |                                                |                           |
|----------------|-------------------|------------------------------------------------|---------------------------|
| • Condensation | • Deposition      | • Calorimeter                                  | • Heat of reaction        |
| • Freezing     | • Heat            | • Joule                                        | • Le Chatlier's principle |
| • Vaporization | • Temperature     | • Specific heat                                | • Equilibrium             |
| • Melting      | • Thermochemistry | • $q = m \cdot c \cdot \Delta T$               | • Activation Energy       |
| • Boiling      | • Exothermic      | • $q = m \cdot \Delta H_{\text{vaporization}}$ | • Catalyst                |
| • Sublimation  | • Endothermic     | • $q = m \cdot \Delta H_{\text{fusion}}$       |                           |

### Study Guide (Use notes, worksheets, and labs in addition to this study guide)

1. According to the collision theory, what 3 things must happen in order for two molecules to react?

*Must collide with each other, Must collide with enough energy, & in the correct orientation*

2. What is the importance of the activation energy in a chemical reaction?

*Amount of energy needed for a reaction to occur.*

3. Using the collision theory, explain if the following would increase/ decrease the rate of reaction:

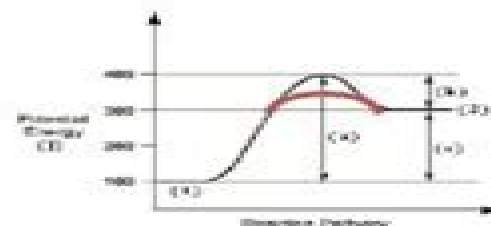


- Increase temp: *Increase, molecules collide more*
- Increase HCl: *Increase, more collisions because more reactants to collide*
- Increase pressure: *Hard to say, wouldn't affect the reactants directly because none of them are gases. One of the products is a gas, so it might cause the gas particles to collide more with the reactants and could slow it down. Tough call.. (this one won't be on the test).*
- Remove  $\text{MgCl}_2$ : *Increase, less products means less chance for them to interfere with reactant collisions*
- Add  $\text{H}_2$ : *Decrease, more products mean they interfere more with reactants and the reactants would collide less.*

4. Identify the following terms with their correct letter from the graphs:

- Reactants- **I**
- Products- **2**
- Activation energy (number/ letter & amount)-  
**A (300 Joules)**
- Is this reaction exothermic or endothermic?  
**Endothermic**

5. Explain how a catalyst affects the activation energy in a chemical reaction. (Also, show on the above graph what a catalyst might do). It lowers the activation energy (see on graph).  
*Speeds it up by lowering the activation energy.*



# Reaction Rates Equilibrium Test Answers

**M Lipman**



## **Reaction Rates Equilibrium Test Answers:**

## Decoding **Reaction Rates Equilibrium Test Answers**: Revealing the Captivating Potential of Verbal Expression

In an era characterized by interconnectedness and an insatiable thirst for knowledge, the captivating potential of verbal expression has emerged as a formidable force. Its ability to evoke sentiments, stimulate introspection, and incite profound transformations is genuinely awe-inspiring. Within the pages of "**Reaction Rates Equilibrium Test Answers**," a mesmerizing literary creation penned by a celebrated wordsmith, readers embark on an enlightening odyssey, unraveling the intricate significance of language and its enduring effect on our lives. In this appraisal, we shall explore the book's central themes, evaluate its distinctive writing style, and gauge its pervasive influence on the hearts and minds of its readership.

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### **Table of Contents Reaction Rates Equilibrium Test Answers**

1. Understanding the eBook Reaction Rates Equilibrium Test Answers
  - The Rise of Digital Reading Reaction Rates Equilibrium Test Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Reaction Rates Equilibrium Test Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Reaction Rates Equilibrium Test Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Reaction Rates Equilibrium Test Answers
  - Personalized Recommendations
  - Reaction Rates Equilibrium Test Answers User Reviews and Ratings
  - Reaction Rates Equilibrium Test Answers and Bestseller Lists

5. Accessing Reaction Rates Equilibrium Test Answers Free and Paid eBooks
  - Reaction Rates Equilibrium Test Answers Public Domain eBooks
  - Reaction Rates Equilibrium Test Answers eBook Subscription Services
  - Reaction Rates Equilibrium Test Answers Budget-Friendly Options
6. Navigating Reaction Rates Equilibrium Test Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Reaction Rates Equilibrium Test Answers Compatibility with Devices
  - Reaction Rates Equilibrium Test Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Reaction Rates Equilibrium Test Answers
  - Highlighting and Note-Taking Reaction Rates Equilibrium Test Answers
  - Interactive Elements Reaction Rates Equilibrium Test Answers
8. Staying Engaged with Reaction Rates Equilibrium Test Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Reaction Rates Equilibrium Test Answers
9. Balancing eBooks and Physical Books Reaction Rates Equilibrium Test Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Reaction Rates Equilibrium Test Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Reaction Rates Equilibrium Test Answers
  - Setting Reading Goals Reaction Rates Equilibrium Test Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Reaction Rates Equilibrium Test Answers
  - Fact-Checking eBook Content of Reaction Rates Equilibrium Test Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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