



Reaction Rates and Equilibrium Study Guide

Topics Covered:

Factors affecting reaction rates
Collision theory of reaction rates
Reaction progress diagrams (activation energy and transition states)
Rate laws and the method of initial rates
Chemical equilibrium
Equilibrium expressions and constants and their meaning
Product-favored vs. Reactant-favored
Shifting equilibrium and Le Chatelier's Principle
Quantitative calculations of equilibrium systems using ICE tables

Practice Questions: [Click here for an answer key](#)

- Describe how each of the following factors would affect the rate of "particle collisions" in your response.
 - Temperature decrease
 - Addition of iodide ions, I^- , as a catalyst
 - Increased reactant concentration
 - Crushing a solid into a powder (with larger surface area)
- The reaction progress diagram shows the energy changes that take place in the reaction of carbon monoxide gas and nitrogen dioxide gas to form carbon dioxide gas.



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Victor M. Corman



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