

Chapter 11: Questions 1-10

Short answer: A long-answer question involving the system's behavior at equilibrium is assigned a maximum of 10 points.

Notes:

Objectives:

Objectives: Complete the following statements for your hypothesis. Consider one of the systems in brackets by highlighting, circling, or placing your answer in a different color.

If the concentration increases, the extent that [equilibrium constant] has been will [increase or decrease] to reach the new equilibrium.

If the concentration increases, the rate of the [forward or reverse] reaction will [increase or decrease] to reach the new equilibrium.

If the concentration increases, the rate of the [forward or reverse] reaction will [increase or decrease] to reach the new equilibrium.

If the concentration increases, the rate of the [forward or reverse] reaction will [increase or decrease] to reach the new equilibrium.

Answer: The student's answer is not expected to be that long. The student's answer should be a short answer that only includes the system, or the concentration, and extent of the reaction.

Short answer: A long-answer question is given with through the student's.

System:	Short answer:
Equilibrium	
Equilibrium	
Equilibrium	

Objectives: Complete your observations and calculations. Record your observations and your calculations below.

Equilibrium constant (K_c) is:

$$K_c = \frac{[P]^p [Q]^q}{[R]^r [S]^s}$$

Objectives: Record all data in your notebook. Record all data in your notebook. Record all data in your notebook.

Reaction Rates And Equilibrium Lab Report Answers

American Rocket Society



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Enjoying the Tune of Expression: An Mental Symphony within **Reaction Rates And Equilibrium Lab Report Answers**

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