

CHAPTER 16 REVIEW***Reaction Energy*****SECTION 2****SHORT ANSWER** Answer the following questions in the space provided.

1. For the following examples, state whether the change in entropy favors the forward or reverse reaction:

- | | |
|-------------------------|--|
| <u>forward reaction</u> | a. $\text{HCl}(l) \rightleftharpoons \text{HCl}(g)$ |
| <u>reverse reaction</u> | b. $\text{C}_6\text{H}_{12}\text{O}_6(aq) \rightleftharpoons \text{C}_6\text{H}_{12}\text{O}_6(s)$ |
| <u>forward reaction</u> | c. $2\text{NH}_3(g) \rightleftharpoons \text{N}_2(g) + 3\text{H}_2(g)$ |
| <u>reverse reaction</u> | d. $3\text{C}_2\text{H}_4(g) \rightleftharpoons \text{C}_6\text{H}_{12}(l)$ |

2. (change in free energy) $\Delta G = (\text{change in enthalpy}) \Delta H - T(\text{change in entropy}) \Delta S$. Write an equation that shows the relationship between enthalpy, ΔH , entropy, ΔS , and free energy, ΔG .

- negative b. For a reaction to occur spontaneously, the sign of ΔG should be _____

3. Consider the following equation: $\text{NH}_3(g) + \text{H}_2\text{O}(l) \rightleftharpoons \text{NH}_4^+(aq) + \text{OH}^-(aq) + \text{energy}$

- true a. The enthalpy factor favors the forward reaction. True or False?
reverse reaction b. The sign of $T\Delta S^\circ$ is negative. This means the entropy factor favors the _____

- c. Given that ΔG° for the above reaction in the forward direction is positive, which term is greater in magnitude and therefore predominates, $T\Delta S$ or ΔH ?

$T(\text{change in } S) \text{ is greater than } (\text{change in } H); (\text{change in } H) \text{ is negative, but } (\Delta G) \text{ is positive, indicating the reaction is not spontaneous. The randomness of the reactants and the temperature of the reaction overshadow the exothermic factor.}$

4. Consider the following equation for the vaporization of water:



- | | |
|-------------------------|--|
| <u>endothermic</u> | a. Is the forward reaction exothermic or endothermic? |
| <u>reverse reaction</u> | b. Does the enthalpy factor favor the forward or reverse reaction? |
| <u>forward reaction</u> | c. Does the entropy factor favor the forward or reverse reaction? |

Reaction Energy Section Answer Key

M Mark



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