

- Q6 When water was added to the anhydrate of $\text{CuSO}_4(x)$, did the reaction proceed in the direction of the reactants or products?

B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of CuCl_2 solution (1)	light blue	light blue
Initial appearance $\text{Cu(OH)}_2(s)$ (2)	clear	clear
Change in appearance (3)	cloudy	clear
Change in appearance (4)	cloudy	clear

- Q7 In test tube 1, did the addition of OH^- cause the reaction to go in the direction of the reactants or products? Explain.
- Q8 In test tube 2, did the addition of H_3O^+ cause the reaction to go in the direction of the reactants or products? Explain.

C. Changing Equilibrium Conditions: Le Châtelier's Principle

C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	$[\text{FeSCN}^{2+}]$ increases or decreases?
1. Reference	red-orange	deeper/lighter	increases/decreases
2. Fe^{3+} added	dark red	deeper	increased
3. SCN^- added	blood red	deeper	increased
4. Fe^{3+} removed	orange	lighter	decreased

Reaction Rates And Equilibrium Lab Answers

S Baum



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Reaction Rates for High-temperature Air with Carbon and Sodium Impurities Mina L. Carnicom, 1968 The values used by a number of investigators for the rate constants of high temperature greater than or equal to 1000 K homogeneous gaseous reactions involving species of the elements nitrogen oxygen carbon and sodium have been compiled and are presented in tabular form Included are reactions involving neutral species charged species free electrons some species in excited electronic or vibrational states and radiative processes Safety-Scale Lab Exp Biochem 2e Spencer L. Seager, Michael R.

Slabaugh, 1994-05 General, Organic, and Biological Chemistry Study Guide and Selected Solutions Karen C.

Timberlake, 2001-11 Keyed to the learning goals in the text this guide is designed to promote active learning through a variety of exercises with answers and mastery exams The guide also contains complete solutions to odd numbered problems

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Instructors Manual to Lab Manual Ralph Petrucci, William Harwood, Geoffrey Herring, 2001 Fossil Energy Update, 1976 Laboratory Manual for Principles of General Chemistry J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment's Report Sheet to provide a space for recording observations and data during the experiment Continued

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