

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a certain compound is 2074°C . Absorption of 4.000 J of energy will result in the decomposition of _____.

- _____ A compound has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If the specific heat capacity is a constant amount for decomposition and is $0.001 \text{ J/g}^\circ\text{C}$, the decomposition from 100°C to 0.001°C is _____.

- The product is a constant from the enthalpy of the product, and the constant for the enthalpy of the product.

- What is the value of ΔH for this reaction?

Reaction Energy Section 1 Answer Key

Clemens Wendtner

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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