

Periodic Trends Worksheet

Directions: Use the trends in electronegativity to answer the questions.

1. Write the following elements by increasing atomic radius: sodium, sulfur, chlorine, oxygen, potassium.

K, Na, O, S, Cl

2. Write the following elements by increasing electronegativity: sodium, oxygen, sulfur, chlorine, potassium.

K, Na, S, Cl, O

3. Which atom has a larger radius: hydrogen or fluorine?

Hydrogen has a larger radius than fluorine because it is further from the nucleus.

4. Which are stronger in the same family: group 1 or group 2 elements?

Group 1 elements are stronger than group 2 elements because they have a larger radius.

5. Evaluate whether the following properties increase or decrease from left to right across the periodic table:

- a. atomic radius: **decreases**
- b. first ionization energy: **increases**
- c. electronegativity: **increases**

6. Which period in the periodic table shows atoms with the greatest atomic radius? Which element has the greatest?

Period 7 shows atoms with the greatest atomic radius. Francium has the greatest atomic radius.

7. Which period in the periodic table shows atoms with the smallest atomic radius? Which element has the smallest?

Period 1 shows atoms with the smallest atomic radius. Hydrogen has the smallest atomic radius.

8. Write the atoms in each group that have the largest atomic radius.

Li	Na	K
Rb	Cs	Fr
Be	Mg	Ca
Sc	Y	La
Ti	Zr	Hf
V	Nb	Ta
Cr	Mo	Tc
Mn	Co	Ni
Fe	Cu	Zn
Ga	In	Tl
Ge	Sn	Pb
As	Sb	Bi
Se	Te	Po
Br	I	At
Hg	Po	At

Al	B	Ga
In	Tl	Bi
Sn	Pb	At
As	Sb	Bi
Te	Po	At
Br	I	At
Hg	Po	At

9. Write the atoms in each group that have the smallest atomic radius.

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Ying Liu



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