

		gerangskik ☒ is om die hoeke tussen hulle te maksimeer. ☒ Strukturele elektronpare is bindingspare plus alleenpare. OF Valensdop-elektronpaarafstotingsmodel is 'n model om die vorms van molekules te voorspel ☒ waarin die elektronpaar vanaf die buitenste dop ☒ van 'n verwysingsatoom rondom hierdie atoom gerangskik is om die afstoting tussen hulle te minimaliseer. ☒ 3 marks/punte	
	1.2.4	The molecule is a tetrahedron ☒ Die molekule is 'n tetraëder ☒ 1 mark/punt	LO2
	1.2.5	 Name trigonal pyramid ☒ Noem trigonale piramidele ☒ 3 marks/punte	LO4
1.3	1.3.1	The electron pair is equally shared between the two Cl atoms ☒ and the resulting Cl₂ molecule has no degree of polarity. ☒ Die elektronpaar word ewe veel gedeel tussen die twee Cl-atome ☒ en die resulterende Cl₂-molekule het geen mate van polariteit nie ☒. 2 marks/punte	LO2
	1.3.2	The tendency ☒ of an atom to pull the bonded pair ☒ of electrons towards itself. ☒ Die neiging ☒ van 'n atoom om die gebinde elektronpare ☒ na homself te trek ☒. 3 marks/punte	LO2
	1.3.3	Electronegativity of Na = 6 Electronegativity of Cl = 3 Difference is 3 ☒ With a difference of greater than 2.1 ☒ electron transfer will take place and the bond ☒ is ionic. ☒ (value of 2.1 does not need to be given for this mark) Elektronegatiwiteit van Na = 6 Elektronegatiwiteit van Cl = 3 Verskil is 3 ☒ Met 'n verskil van meer as 2.1 ☒ sal elektroneoordrag plaasvind en die binding is ionies. ☒ (waarde van 2.1 hoef nie vir hierdie punt gegee te word nie) 4 marks/punte	LO2

Physical Science Control Test 2014 Grade 1

J Dewey



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