



THE AMERICAN SAILING ASSOCIATION'S
SAILING
MADE EASY

THE OFFICIAL MANUAL FOR THE ASA BASIC KEELBOAT SAILING COURSE

Sailing Made Easy

Christian G. Meyer



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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g})$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction $\text{C}_2\text{H}_4(\text{g}) + 2 \text{C}_2\text{H}_6(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + 2 \text{C}_2\text{H}_6(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ/mole}$
 $\text{C}_2\text{H}_6(\text{g}) + 7/2 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ/mole}$
 $\text{H}_2(\text{g}) + 1/2 \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$

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web hess s law states that the total enthalpy change for the reaction will be the sum of all those changes no matter how many different steps or stages in the reaction there are cohen 2016 the equations for the reactions in the experiment done are as follows $\text{NaOH} + \text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O}$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of NaOH combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}}$

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}^+(\text{aq}) + \text{OH}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $2\text{NaOH}(\text{s}) + \text{H}^+(\text{aq}) + \text{Cl}^-(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l}) + \text{Na}^+(\text{aq}) + \text{Cl}^-(\text{aq})$ $\text{H}_2\text{O}(\text{l})$