

Balancing redox reactions under Basic Conditions



- Let's balance the reduction one first
 - for every Oxygen add a water on the other side
 - For every hydrogen add a H^+ to the other side
 - Each H^+ will react with an OH^- on both sides (*whatever you do to one side...you do the same to the other*)
 - H^+ and OH^- make water
 - cancel the waters
 - Balance the imbalance of charge with electrons (-1 vs. -7)
- $\text{ClO}_3^- \rightarrow \text{Cl}^-$
 - $\text{ClO}_3^- \rightarrow \text{Cl}^- + 3\text{H}_2\text{O}$
 - $6\text{H}^+ + \text{ClO}_3^- \rightarrow \text{Cl}^- + 3\text{H}_2\text{O}$
 - $6\text{OH}^- + 6\text{H}^+ + \text{ClO}_3^- \rightarrow \text{Cl}^- + 3\text{H}_2\text{O} + 6\text{OH}^-$
 - $6\text{H}_2\text{O} + \text{ClO}_3^- \rightarrow \text{Cl}^- + 3\text{H}_2\text{O} + 6\text{OH}^-$
 - $3\text{H}_2\text{O} + \text{ClO}_3^- \rightarrow \text{Cl}^- + 6\text{OH}^-$
 - $6e^- + 3\text{H}_2\text{O} + \text{ClO}_3^- \rightarrow \text{Cl}^- + 6\text{OH}^-$

Pogil Balancing Redox Equations

Lingsheng Yao



Pogil Balancing Redox Equations:

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