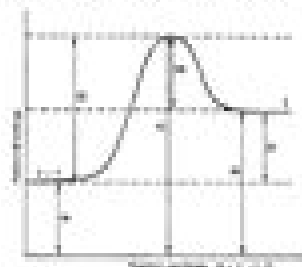
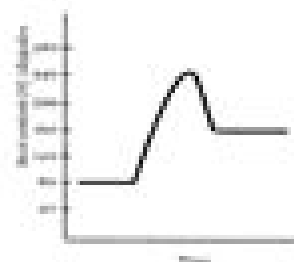


Potential Energy Diagram Worksheet



1. Which of the letters a-f in the diagram represents the potential energy of the products? e
2. Which letter indicates the potential energy of the activated complex? c
3. Which letter indicates the potential energy of the reactants? a
4. Which letter indicates the activation energy? b
5. Which letter indicates the heat of reaction? f
6. Is the reaction endothermic or exothermic? Endo
7. Which letter indicates the activation energy of the reverse reaction? a
8. Which letter indicates the heat of reaction of the reverse reaction? f
9. Is the reverse reaction endothermic or exothermic? EXO



1. The heat content of the reactants of the forward reaction is about 80 kilojoules.
2. The heat content of the products of the forward reaction is about 20 kilojoules.
3. The heat content of the activated complex of the forward reaction is about 160 kilojoules.
4. The activation energy of the forward reaction is about 80 kilojoules.
5. The heat of reaction (ΔH) of the forward reaction is about -80 kilojoules.
6. The forward reaction is Exo (endothermic or exothermic).
7. The heat content of the reactants of the reverse reaction is about 20 kilojoules.
8. The heat content of the products of the reverse reaction is about 80 kilojoules.
9. The heat content of the activated complex of the reverse reaction is about 160 kilojoules.
10. The activation energy of the reverse reaction is about 80 kilojoules.
11. The heat of reaction (ΔH) of the reverse reaction is about +80 kilojoules.
12. The reverse reaction is EXO (endothermic or exothermic).

Key Energy 10.1: Reaction Rates

1. Chemical reactions occur when reactants collide. For what reasons may a collision fail to produce a chemical reaction?

Not enough energy or wrong angle

2. If every collision between reactants lead to a reaction, what determines the rate at which the reaction occurs?

of effective collisions

3. What is the activation energy of a reaction, and how is this energy related to the activated complex of the reaction?

Ea is the energy needed to form the A.C.

4. What happens when a catalyst is used in a reaction?

Catalysts lower the Ea, thus speeding up reactions

5. Name 4 things that will speed up or slow down a chemical reaction.

- ΔT Temp
- catalyst
- C 3
- ΔP Press.

6. Draw an energy diagram for a reaction. (label the axis)

Potential energy of reactants = 350 kJ/mole
Activation energy = 200 kJ/mole
Potential energy of products = 150 kJ/mole



Time

7. Is the reaction in #6 exothermic or endothermic? Explain.

Exo, because P has less E than R

8. How could you lower the activation energy for the reaction in #6?

Use a catalyst.

Pogil Answer Key Intermolecular

Clifford Lane



Pogil Answer Key Intermolecular:

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