

Calculating pH

How will it be used for the concentration differences here?

Why?

As buffers and other substances control pH in organisms, it is important to understand the ability of buffers to resist change. This property is called **buffering capacity** and is usually expressed as millimoles per litre of buffer that will change pH by one unit. Thus, a buffer with a buffering capacity of 10 mmol/L will resist 10 mmol/L of strong acid or base. The buffering capacity of a buffer is related to the concentration of the buffer components. A buffer with a higher concentration of buffer components will have a higher buffering capacity.

Table 1 – Ion Concentrations for Acids and Bases

Acid	Substance	Acidic, Basic or Neutral?	$[H^+]$ (M)	$[OH^-]$ (M)	pH or pOH
1	0.1 M HCl (strong)	Acidic	1.0×10^{-1} M	1.0×10^{-14} M	1.0
2	0.001 M HCl (strong)	Acidic	1.0×10^{-3} M	1.0×10^{-11} M	3.0
3	0.00001 M HCl (strong)	Acidic	1.0×10^{-5} M	1.0×10^{-9} M	5.0
4	0.000001 M HCl (strong)	Acidic	1.0×10^{-6} M	1.0×10^{-8} M	6.0
5	0.0000001 M HCl (strong)	Acidic	1.0×10^{-7} M	1.0×10^{-7} M	7.0
6	0.00000001 M HCl (strong)	Acidic	1.0×10^{-8} M	1.0×10^{-6} M	8.0
7	0.000000001 M HCl (strong)	Acidic	1.0×10^{-9} M	1.0×10^{-5} M	9.0
8	0.0000000001 M HCl (strong)	Acidic	1.0×10^{-10} M	1.0×10^{-4} M	10.0
9	0.00000000001 M HCl (strong)	Acidic	1.0×10^{-11} M	1.0×10^{-3} M	11.0
10	0.000000000001 M HCl (strong)	Acidic	1.0×10^{-12} M	1.0×10^{-2} M	12.0
11	0.0000000000001 M HCl (strong)	Acidic	1.0×10^{-13} M	1.0×10^{-1} M	13.0

1. What does the number 13.0 pH indicate? (strong)

It indicates that the solution is very acidic.

2. For Solution 1, which one has a higher concentration, hydrochloric acid or hydroxide ions?

Hydrochloric acid.

3. How did you determine the concentration of hydrochloric acid ions in Solution 1 or Solution 2? How did you determine the concentration of the acid?

From the pH value, I determined the concentration of the acid. I used the formula $pH = -\log[H^+]$ to find the concentration of the acid.

4. How did you determine the concentration of hydroxide ions in Solution 1 or Solution 2? How did you determine the concentration of the base?

From the pH value, I determined the concentration of the acid. I used the formula $pH = -\log[H^+]$ to find the concentration of the acid.

Pogil Calculating Ph Answers

Léo Pomar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the name bar.

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