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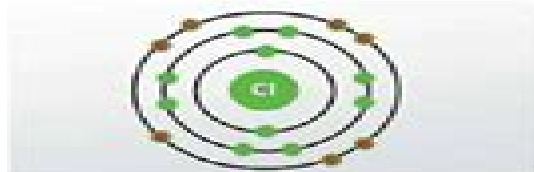
How are electrons shared to create covalently bonded molecules?

Read This!

Why? Covalent bonds result from a *sharing* of electrons by two or more atoms (usually nonmetals). Lewis theory (Gilbert Newton Lewis, 1875-1946) focuses on the valence electrons, since the outermost electrons are the ones that are highest in energy and farthest from the nucleus, and are therefore the ones that are most exposed to other atoms when bonds form. Lewis dot diagrams for elements are a handy way of picturing valence electrons, and especially, which electrons are available to be shared in covalent bonds.

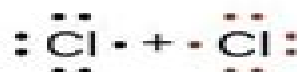
Bohr Model: Chlorine

7 valence electrons on outermost energy level



Lewis Dot diagram: Chlorine

7 valence electrons



chlorine
atoms



chlorine
molecule



Lewis Dot Diagrams Showing Valence Electrons for Individual Elements

1	2		3	4	5	6	7	8
$\cdot\text{H}$								He
$\cdot\text{Li}$	$\cdot\text{Be}\cdot$		$\cdot\text{B}\cdot$	$\cdot\text{C}\cdot$	$\cdot\text{N}\cdot$	$\cdot\text{O}\cdot$	$\cdot\text{F}\cdot$	$\cdot\text{Ne}\cdot$
$\cdot\text{Na}$	$\cdot\text{Mg}\cdot$	Transition metals	$\cdot\text{Al}\cdot$	$\cdot\text{Si}\cdot$	$\cdot\text{P}\cdot$	$\cdot\text{S}\cdot$	$\cdot\text{Cl}\cdot$	$\cdot\text{Ar}\cdot$

Atoms are grouped together by number of valence electrons; all atoms in Group 1 have 1 valence electron.

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WJ Hussar

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

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