

Name: _____

Neutralization Reactions Worksheet

Reading: Cycle 6 Chemistry III Lesson 4

Neutralization Reaction: Acid + Base $\rightarrow$ Salt + Water

H⁺ from acid and OH⁻ from base combine to form H₂O. Leftovers form 'salt'.

Remember in salt, write metal first! KCl not ClK!

Example 1: $\text{HI} + \text{NaOH} \rightarrow \text{H}_2\text{O} + \text{NaI}$

H and OH combine to form water. One atom of I & one of Na leftover. Na is a metal so write NaI (not INa)

Example 2: $\text{HF} + \text{Ca(OH)}_2 \rightarrow \text{ } + \text{ }$

For complete neutralization, must have equal number of H and OH. Therefore we need 2 HF for 1 Ca(OH)₂

$2\text{HF} + 1\text{Ca(OH)}_2 \rightarrow \text{ } + \text{ }$

Two H and two OH make two waters. Think of procs! Two boys and two girls make two couples.

$2\text{HF} + 1\text{Ca(OH)}_2 \rightarrow 2\text{H}_2\text{O} + \text{ }$

What's left? 2 F atoms and 1 Ca atom form the "salt". Remember that Ca is the metal, so it goes first.

$2\text{HF} + 1\text{Ca(OH)}_2 \rightarrow 2\text{H}_2\text{O} + \text{CaF}_2$

Safety check: is CaF₂ a good formula? P.T. tells you ions are Ca⁺² and F⁻, so total charge is 0 like it should be.

1. Complete & balance these neutralization reactions. Remember that H₂O = HOH!



2. Complete and balance the following equations representing neutralization reactions:



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Albert A Gayle



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