

Empirical and Molecular Formulas Practice

Read the question carefully for the data provided. Answers must be shown for credit.

Find empirical formula:

1. A compound is found to contain 40.0 g of carbon and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{40.0\text{g}} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 3.33 \text{ mol C}$$

$$\text{H}_{6.67\text{g}} \times \frac{1 \text{ mol H}}{1.008 \text{ g H}} = 6.62 \text{ mol H}$$

C₃H₅

2. A compound is found to contain 20.0 g of carbon, 26.7 g of oxygen, and 6.67 g of hydrogen. Find the empirical formula.

$$\text{C}_{20.0\text{g}} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C}$$

$$\text{O}_{26.7\text{g}} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 1.67 \text{ mol O}$$



$$\text{C}_1\text{O}_1\text{H}_4 = \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.67 \text{ mol C} \left(\frac{1 \text{ mol C}}{1.67 \text{ mol C}} = 1 \right)$$

3. A 10.0 g sample of a compound contains 2.00 g carbon, 1.33 g hydrogen, and 6.67 g oxygen. Find the empirical formula.

$$\text{C}_{2.00\text{g}} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.167 \text{ mol C}$$



$$\text{O}_{6.67\text{g}} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.417 \text{ mol O} \left(\frac{0.167 \text{ mol C}}{0.167 \text{ mol C}} = 1 \right)$$

$$\text{C}_1\text{H}_8\text{O}_2 = \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 0.167 \text{ mol C} \left(\frac{1.32 \text{ mol H}}{0.167 \text{ mol C}} = 8 \right)$$

4. A 20.00 g sample of an unknown compound contains 12.00 g carbon, 8.00 g hydrogen, and 0.00 g oxygen. Find the empirical formula.

$$\text{C}_{12.00\text{g}} \times \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C}$$



$$\text{O}_{0.00\text{g}} \times \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 0.00 \text{ mol O}$$

$$\text{C}_1\text{H}_8\text{O}_0 = \frac{1 \text{ mol C}}{12.01 \text{ g C}} = 1.00 \text{ mol C} \left(\frac{7.94 \text{ mol H}}{1.00 \text{ mol C}} = 8 \right)$$

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Hussin A.Rothana



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