

Gas Variables

How are the variables that describe a gas related?

Why?

Imagine buying a balloon bouquet at a party store. How will the helium gas in the bouquet behave if you carry it outside on a hot summer day? How will it behave if you carry it outside during a snowstorm? What happens if the balloons are made of latex, which can stretch? What happens if the balloons are made of Mylar[®], which cannot stretch? What if you add just a small amount of gas to each balloon? What if you add a lot of gas? In this activity, you will explore four variables that quantify gases—pressure (P), volume (V), temperature (T), and moles (n) of gas. These four variables can be related mathematically so that predictions about gas behavior can be made.

Model 1 – Gases in a Nonflexible Container

Experiment A (Adding more gas)



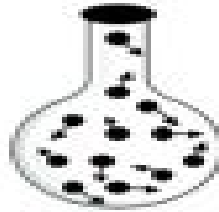
A1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm
Temperature = 200 K



A2

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm
Temperature = 200 K



A3

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm
Temperature = 200 K

Experiment B (Heating the gas)



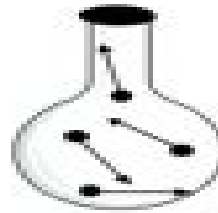
B1

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 1 atm
Temperature = 200 K



B2

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 2 atm
Temperature = 400 K



B3

Volume = 1 unit
External pressure = 1 atm
Internal pressure = 3 atm
Temperature = 600 K

**Note: Volume in this model is recorded in units rather than liters because 4 molecules of gas at the conditions given would occupy a very small space ($\sim 1 \times 10^{-22}$ gL). The particles shown here are much larger compared to the space between them than actual gas particles.*

Pogil Activities For High School Chemistry Gas Variables Answers

RC Schank

A decorative graphic consisting of a light blue horizontal bar with a rounded right end, and a red circular shape partially visible behind it.

Pogil Activities For High School Chemistry Gas Variables Answers:

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Table of Contents Pogil Activities For High School Chemistry Gas Variables Answers

1. Understanding the eBook Pogil Activities For High School Chemistry Gas Variables Answers
 - The Rise of Digital Reading Pogil Activities For High School Chemistry Gas Variables Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Activities For High School Chemistry Gas Variables Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Activities For High School Chemistry Gas Variables Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Activities For High School Chemistry Gas Variables Answers
 - Personalized Recommendations
 - Pogil Activities For High School Chemistry Gas Variables Answers User Reviews and Ratings
 - Pogil Activities For High School Chemistry Gas Variables Answers and Bestseller Lists
5. Accessing Pogil Activities For High School Chemistry Gas Variables Answers Free and Paid eBooks
 - Pogil Activities For High School Chemistry Gas Variables Answers Public Domain eBooks
 - Pogil Activities For High School Chemistry Gas Variables Answers eBook Subscription Services
 - Pogil Activities For High School Chemistry Gas Variables Answers Budget-Friendly Options

6. Navigating Pogil Activities For High School Chemistry Gas Variables Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Activities For High School Chemistry Gas Variables Answers Compatibility with Devices
 - Pogil Activities For High School Chemistry Gas Variables Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Activities For High School Chemistry Gas Variables Answers
 - Highlighting and Note-Taking Pogil Activities For High School Chemistry Gas Variables Answers
 - Interactive Elements Pogil Activities For High School Chemistry Gas Variables Answers
8. Staying Engaged with Pogil Activities For High School Chemistry Gas Variables Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Activities For High School Chemistry Gas Variables Answers
9. Balancing eBooks and Physical Books Pogil Activities For High School Chemistry Gas Variables Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Activities For High School Chemistry Gas Variables Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Activities For High School Chemistry Gas Variables Answers
 - Setting Reading Goals Pogil Activities For High School Chemistry Gas Variables Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Activities For High School Chemistry Gas Variables Answers
 - Fact-Checking eBook Content of Pogil Activities For High School Chemistry Gas Variables Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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