

Free Energy

What does this energy mean? Does this always tell us everything about?

Read this: Although thermodynamics is a science that is always true in the sense of that big issue, in practice, if the flow of energy is not understood in full, we have an incomplete look at systems in that there are other things in the environment with which they interact. (part of the idea of Conservation of Energy) So why consider these things but enough energy to go around?

Model 1 – Spontaneous Processes

	Process Description	Δ Energy in Surroundings (J/K)	Δ Energy in System (J/K)	Spontaneous?
a.	Heat flows from a warmer object to a cooler one	> 0	Decreasing	Yes
b.	Rolling an object down a hill. Includes an increase in the kinetic energy of the object	> 0	Decreasing	Yes
c.	Freezing water. Includes an increase in the order of the water molecules	Decreasing (negative)	Increasing (positive)	Yes
d.	$\text{H}_2\text{O}(l) + \text{H}_2\text{O}(l) \rightarrow \text{H}_3\text{O}^+(aq) + \text{OH}^-(aq)$	Heat transfer	Decreasing	Yes
e.	$\text{H}_2\text{O}(l) + \text{H}_2\text{O}(l) \rightarrow \text{H}_2\text{O}_2(l) + \text{H}_2(g)$	Endothermic	Heat transfer	No
f.	$\text{H}_2\text{O}(l) \rightarrow \text{H}_2(g) + \text{O}_2(g)$	Endothermic	Heat transfer	No
g.	Water is hot	Heat transfer	Heat transfer	Water at 70°C can become at 75°C but
h.	Equal masses (25°C) of hot water (50 g) and cold water (50 g) are mixed	Endothermic	Decreasing	Water at 25°C can become at 25°C but

1. Complete Model 1.

2. Spontaneous means we would be describing the surrounding, it is not the the surrounding?

Free Energy or Gibbs Free Energy

3. What is the equation for the change in free energy?

4. What does ΔG tell us about the spontaneity of a process? Does it tell us about the energy in the surroundings?

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M Tight



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