Naming Ionic Compounds

What are the structural units that make up ionic compounds and how are they named?

Why?

When working in chemistry, it is often convenient to write a chemical in symbols. For example we might write down the substance table salt as NaCl. In talking about chemistry however, it is a bit tacky to say "en-ay see-ell" when we want to refer to a substance. Also, in formal writing we should use the name of the compound rather than its symbols. Therefore we need to learn how to say the proper names of ionic substances.

Model 1 – Ion Charges for Selected Elements

H:		_								
E3°	Be ³ *							Nº-	O3-	E1-
Nat	Mg ²	Transition elements				Al ³		Pa-	S	CP
601	Cas	Fe ² *	NP-	Cu*	Zn3+					Bett
Rb*	Selv			Agr			Sm ²⁻ Sm ⁴⁻			127
	Ba2-				Hg ₂ .		Pb**			

- Based on the information in Model 1:
 - a. Identify three elements that form only one cation. Posty in I on

Nat, 5-2+, A12+ (Any box with one element works.)

- 6. Identify three elements that form only one anion.

 6. Identify three elements that form more than one cation.

 6. Identify three elements that form more than one cation.

 Fet of fet, Nitand Nit, Cut and Cut.

 A. In what region of the periodic table are these "multiple ion" elements usually located?

 Trans. From elements or of block

- Amiona -

- Carions

Consider the ions of potassium (K) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of potassium (K) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

3. Consider the ions of iron (Fe) and sulfur (S). Write chemical formulas for all possible ionic compounds involving these ions, using the simplest ratio(s) of iron (Fe) and sulfur (S). Keep in mind that the sum of the charges in an ionic compound must equal zero.

Fe and 52 (Fe5) and 2Fe3+ and 353 (Fe253)

Namine Ionic Compounds

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