

Model 2 – Prefixes and Suffixes

Prefix	Numerical Value
mono-	1
di-	2
tri-	3
tetra-	4
penta-	5
hexa-	6
hepta-	7
octa-	8
nona-	9
deca-	10

Molecular Formula	Name of Compound
BCl_3	Boron trichloride
SF_6	Sulfur hexafluoride
IF_7	Iodine heptafluoride
NI_3	Nitrogen triiodide
N_2O_4	Dinitrogen tetroxide
Cl_2O	Dichlorine monoxide
P_4O_{10}	Tetraphosphorus decoxide
B_5H_9	Pentaboron nonahydride
Br_3O_8	Tribromine octoxide
ClF	Chlorine monofluoride

- Examine the prefixes in Model 2. Fill in the numerical value that corresponds to each prefix.
- What suffix (ending) do all the compound names in Model 2 have in common?

They all end in "ide".

- Carefully examine the names of the compounds in Model 2. When is a prefix NOT used in front of the name of an element?

*When the first element is one. (Examples: IF_7 - iodine heptafluoride)
 NI_3 - nitrogen triiodide)*

- Consider the compound NO .

- Which element, nitrogen or oxygen, would require a prefix in the molecule name? Explain your answer.

Oxygen would require a prefix of "mono" because it is the second element and there is one oxygen. Nitrogen does not require a prefix because it is the first element and there is one nitrogen.

- Name the molecule NO .

*Nitrogen monoxide (Notice, the "o" of mono was dropped)
 ↳ not monoxide*

- Find two compounds in Model 2 that contain a subscript of "4" in their molecular formula.

- List the formulas and names for the two compounds.

N_2O_4 - dinitrogen tetroxide || P_4O_{10} - Tetraphosphorus decoxide

- What is different about the spelling of the prefix meaning "four" in these two names?

The "a" has been dropped from the end of the prefix tetra.

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