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POGIL-The Mole Concept (How can atoms be counted?)

Why?

Consider the following equation for a chemical reaction: $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$

This can be interpreted as two molecules of hydrogen and one molecule of oxygen combining to form two water molecules. But how often do chemists limit their reactions to one or two molecules? Usually a reaction is done with an unimaginable number of molecules. How then do chemists know they have the right mix? The molecules need to be quickly counted! How do we count molecules? The answer is the unit called the **mole**. Weighing is often easier than counting!

Model 1: The Balanced Equation

How can the coefficients in a chemical equation be interpreted?

Consider the reaction:



Key Questions

1. What are the coefficients for each of the following substances in the reaction?

Read This!

Coefficients tell you the relative amounts of products and reactants. A balanced chemical equation can be interpreted in terms of different quantities, including:

- Numbers of particles (atoms, molecules, or formula units)

PARTICLES

Atoms



(if you have a single element)

- Mass
- Volume
- **Moles**

Molecules



(if you have a covalent compound)

Formula Units



(if you have an ionic compound)



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