

Intermolecular Forces Notes

Section 2

What is the Lewis Dot Structure for Carbon? Based on this Lewis Dot Structure, how many electrons can Carbon bond with?



Carbon only has two valence electrons that are represented available for bonding. So, carbon should only be able to bond with TWO atoms.

So how do we explain the chemical CH_4 ? Two theories: (1) Valence Bond Theory (2) Molecular Orbital Theory

Valence Bond Theory (VBT Theory)

It is it carbon can hybridize, meaning they mix together.

Orbital Diagram of Carbon:



Here, all 2s orbitals and 2 p orbitals have hybridized. The orbitals sp hybridization.

Types of hybridization: (1) sp (2) sp² (3) sp³

"A "connection" can be a BOND (line) or a LONE PAIR (2 dots)

Practice Problem:

What type of hybridization occurs in the molecule PF_3 ?



Since the molecule has THREE connections around P, it's sp³ hybridization.



Methane: CH_4

(2) Methylamine: CH_3NH_2

Practice Problem:

What type of hybridization, if any, occurs in the following molecules?

(1) BF_3

(2) BCl_3



sp² around B



sp² around B

Valence Bond Theory includes a trick

A single bond is called a sigma bond. Its symbol is σ

A double bond is 2 sigma & 2 pi bonds. Its symbol is π

A triple bond contains 1 sigma and 2 pi bonds.



Practice Problem:

How many sigma bonds are in this molecule of acetone? How many pi bonds?

3 sigma
1 pi



Practice Problem:

Use valence bond theory to describe the bonding in N_2 .



Each nitrogen has 4 hybridization.
There is 1 sigma & 2 pi bonds.

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Table of Contents Pogil Intermolecular Forces Answer Key Stbedes

1. Understanding the eBook Pogil Intermolecular Forces Answer Key Stbedes
 - The Rise of Digital Reading Pogil Intermolecular Forces Answer Key Stbedes
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Intermolecular Forces Answer Key Stbedes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Intermolecular Forces Answer Key Stbedes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Intermolecular Forces Answer Key Stbedes
 - Personalized Recommendations
 - Pogil Intermolecular Forces Answer Key Stbedes User Reviews and Ratings
 - Pogil Intermolecular Forces Answer Key Stbedes and Bestseller Lists
5. Accessing Pogil Intermolecular Forces Answer Key Stbedes Free and Paid eBooks

- Pogil Intermolecular Forces Answer Key Stbedes Public Domain eBooks
 - Pogil Intermolecular Forces Answer Key Stbedes eBook Subscription Services
 - Pogil Intermolecular Forces Answer Key Stbedes Budget-Friendly Options
6. Navigating Pogil Intermolecular Forces Answer Key Stbedes eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Intermolecular Forces Answer Key Stbedes Compatibility with Devices
 - Pogil Intermolecular Forces Answer Key Stbedes Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Intermolecular Forces Answer Key Stbedes
 - Highlighting and Note-Taking Pogil Intermolecular Forces Answer Key Stbedes
 - Interactive Elements Pogil Intermolecular Forces Answer Key Stbedes
 8. Staying Engaged with Pogil Intermolecular Forces Answer Key Stbedes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Intermolecular Forces Answer Key Stbedes
 9. Balancing eBooks and Physical Books Pogil Intermolecular Forces Answer Key Stbedes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Intermolecular Forces Answer Key Stbedes
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Pogil Intermolecular Forces Answer Key Stbedes
 - Setting Reading Goals Pogil Intermolecular Forces Answer Key Stbedes
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Pogil Intermolecular Forces Answer Key Stbedes
 - Fact-Checking eBook Content of Pogil Intermolecular Forces Answer Key Stbedes
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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