

Phases and Intermolecular Forces

Each of the following sets of molecules are similar in either Electron Pair Geometry or Molecular Geometry. Even when the shapes are the same the polarity may vary. Identify shape and polarity for each structure, and then consider the questions posed.

1. There are several ways to be linear. How are they different (linear Electron Geometry, linear Molecular Geometry, both linear EG and MG)?



- a. Explain the difference in polarity for HCN and CO_2 ?

2. There are two ways to be bent. How are they different?

They're different by shape + angle of molecule

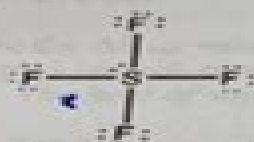


- a. What would be the formal charges, if SO_2 had a 1 single/1 double bond arrangement?

- b. Are these polar or nonpolar molecules?

- c. What shape is H_2O ?

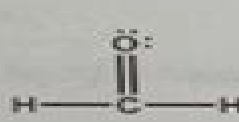
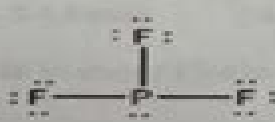
3. Tetrahedral or not?



- a. Why are two molecules nonpolar and one molecule polar?

SF_4 is polar being that electrons are throwing off electronegativity.

4. What makes trigonal pyramidal and trigonal planar molecules different? *Causing change in shape*



- a. Why are two of the molecules polar and one molecule nonpolar?

Pogil Intermolecular Forces And Boiling Points Answers

Katrin Zwirgmaier



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