

Naming Molecular Compounds

Why?

How are the chemical formula and name of a molecular compound related?

When you began chemistry class this year, you probably already knew that the chemical formula for carbon dioxide was CO_2 . Today you will find out why CO_2 is named that way. Naming chemical compounds correctly is of paramount importance. The slight difference between the names carbon monoxide (CO , a poisonous, deadly gas) and carbon dioxide (CO_2 , a greenhouse gas that we exhale when we breathe out) can be the difference between life and death! In this activity you will learn the naming system for molecular compounds.

Model 1 – Molecular Compounds

Molecular Formula	Number of Atoms of First	Number of Atoms of Second	Name of Compound
ClF	1	1	Chlorine monofluoride
ClF_5	1	5	Chlorine pentafluoride
CO	1	1	Carbon monoxide
CO_2	1	5	Carbon dioxide
Cl_2O	2	1	Dichlorine monoxide
PCl_5	1	5	Phosphorus pentachloride
N_2O_5	2	5	Dinitrogen pentoxide

- Fill in the table to indicate the number of atoms of each type in the molecular formula.
- Examine the molecular formulas given in Model 1 for various molecular compounds.
 - How many different *elements* are present in each compound shown?
Two
 - Do the compounds combine metals with metals, metals with nonmetals, or nonmetals with nonmetals?
Non-metals with non-metals
 - Based on your answer to b, what type of bonding must be involved in molecular compounds?
Covalent bonds
- Find all of the compounds in Model 1 that have chlorine and fluorine in them. Explain why the name "chlorine fluoride" is not sufficient to identify a specific compound.
The number of fluorine atoms is different in the two compounds, but name does not reflect the difference
- Assuming that the name of the compound gives a clue to its molecular formula, predict how many atoms each of these prefixes indicates, and provide two examples.

Pogil Chemistry Answer Key Naming Molecular Compounds

JR Anderson



Pogil Chemistry Answer Key Naming Molecular Compounds:

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