

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are (broken/formed) there is a positive energy change.
 - Breaking bonds is (endothermic/exothermic).
 - When bonds are (broken/formed) there is a negative energy change.
 - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
 - How do the changes in energy for the reverse reactions compare?
The changes in energy for these reactions are the same, but opposite in sign.
 - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.
If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

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Table of Contents Pogil Bond Energy Key

1. Understanding the eBook Pogil Bond Energy Key
 - The Rise of Digital Reading Pogil Bond Energy Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Pogil Bond Energy Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Pogil Bond Energy Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Pogil Bond Energy Key
 - Personalized Recommendations
 - Pogil Bond Energy Key User Reviews and Ratings
 - Pogil Bond Energy Key and Bestseller Lists

5. Accessing Pogil Bond Energy Key Free and Paid eBooks
 - Pogil Bond Energy Key Public Domain eBooks
 - Pogil Bond Energy Key eBook Subscription Services
 - Pogil Bond Energy Key Budget-Friendly Options
6. Navigating Pogil Bond Energy Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Pogil Bond Energy Key Compatibility with Devices
 - Pogil Bond Energy Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Pogil Bond Energy Key
 - Highlighting and Note-Taking Pogil Bond Energy Key
 - Interactive Elements Pogil Bond Energy Key
8. Staying Engaged with Pogil Bond Energy Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Pogil Bond Energy Key
9. Balancing eBooks and Physical Books Pogil Bond Energy Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Pogil Bond Energy Key
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Pogil Bond Energy Key
 - Setting Reading Goals Pogil Bond Energy Key
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Pogil Bond Energy Key
 - Fact-Checking eBook Content of Pogil Bond Energy Key
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

-
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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